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## Electrocatalytic gas sensor with reference layer

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### Abstract

This paper presents studies of gas sensors prepared in ceramic technology with Nasicon as a solid electrolyte. Sensors work in the voltammetric mode thus based on the excitation of a sensor with a periodic potential signal while current response is recorded. The main aim is to investigate a Bi<sub>8</sub>Nb<sub>2</sub>O<sub>17</sub> reference layer influence on sensor properties. Sensors I-V characteristics in different concentration of nitrogen dioxide have been measured. Three different constructions of the sensor are compared.

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### 1. Introduction

In recent years electrochemical gas sensors based on solid state electrolytes have been intensively developed. They are easy to obtain, use and relative durable [1]. Nasicon is one of the most promising materials, which have been used in construction of gas sensors based on solid electrolytes. Most of these devices work in potentiometric or amperometric mode. However, some works are dedicated to sensors working in electrocatalytic mode [2, 3]. Reported results indicate that such sensors have enhanced selectivity [3]. Principle of operation of such sensors is based on the galvanic cell excitation with a periodic potential signal, while current response is recorded. When a voltage is applied to sensors electrodes an oxidation or a reduction of chemical species occurs. This results in a unique current-voltage (I-V) characteristic for each type and concentration of measured gas mixture. As a consequence gas concentration determination is possible.

Electrocatalytic sensors usually employs simple structure of a solid electrolyte and two metal electrodes [3, 4]. In this study electrocatalytic sensor with additional, reference layer is proposed. Our main goal was to check if presence

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of  $\text{Bi}_8\text{Nb}_2\text{O}_{17}$  reference layer has any beneficial effect on sensor properties in case of electrocatalytic nitrogen dioxide sensor. In case of potentiometric gas sensor with such layer an improvement of sensor properties have been reported [5]. Properties of the sensors with and without this layer are compared.

## 2. Experimental

### 2.1. Sensor preparation

Nasicon pellets were prepared using the sol-gel method [6]. As a substrates  $\text{NH}_4\text{H}_2\text{PO}_4$ ,  $\text{Na}_2\text{SiO}_3 \cdot 9\text{H}_2\text{O}$  and  $\text{ZrO}(\text{NO}_3)_2 \cdot x\text{H}_2\text{O}$  were used. Aqueous solution of  $\text{ZrO}(\text{NO}_3)_2$  was mixed with aqueous solutions of  $\text{NH}_4\text{H}_2\text{PO}_4$  and  $\text{Na}_2\text{SiO}_3 \cdot 9\text{H}_2\text{O}$  (molar ratio of components was 2:1:2) to form a sol. In order to obtain a gel, sol was heated at  $100^\circ\text{C}$  for 12 hours and continuously dried at  $100^\circ\text{C}$  to form xerogel. First calcination at  $750^\circ\text{C}$  for 2 hours and afterwards sintering at  $1000^\circ\text{C}$  for 12 hours were performed. Reference electrode material ( $\text{Bi}_8\text{Nb}_2\text{O}_{17}$ ) was prepared from  $\text{Bi}_2\text{O}_3$  and  $\text{Nb}_2\text{O}_5$ . These oxides were mixed at the molar ratio 4:1 ( $\text{Bi}_2\text{O}_3:\text{Nb}_2\text{O}_5$ ). Calcination of oxides powders at  $750^\circ\text{C}$  for 5 hours and sintering at  $950^\circ\text{C}$  for 10 hours were performed.

One side of a Nasicon pellets was painted with Pt paste and then sintered at  $900^\circ\text{C}$ . On the opposite side of pellet a layer of  $\text{Bi}_8\text{Nb}_2\text{O}_{17}$  and Pt mixture was obtained in similar manner. Two kinds of sensor with reference layer were investigated, one with both electrodes exposed to the test gas mixture and other with Pt+ $\text{Bi}_8\text{Nb}_2\text{O}_{17}$  electrode coated with dielectric sealing. Properties of these two sensors were compared with results obtained for the symmetrical sensor without the reference electrode. Investigated sensors structures are presented in Fig. 1 and referred later as a sensor 1, sensor 2 and sensor 3.

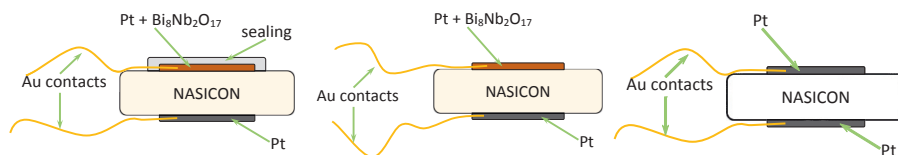


Fig. 1. Sensor 1 with additional layer sealed (a), sensor 2 with additional layer unsealed (b) and sensor 3 with symmetrical structure (c)

### 2.2. Sensor properties measurements

Measurement stand is presented in the Fig. 2. The measurements were performed using the electrochemical interface SI 1287 and a PC with suitable software for system control and data acquisition. Linearly changing voltage of symmetrical triangular shape (range from 5 V to  $-5$  V) was used. Temperature was fixed to  $300^\circ\text{C}$ .

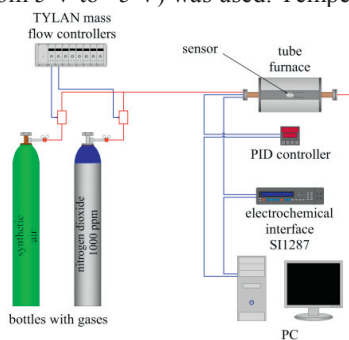


Fig. 2. Measurement stand

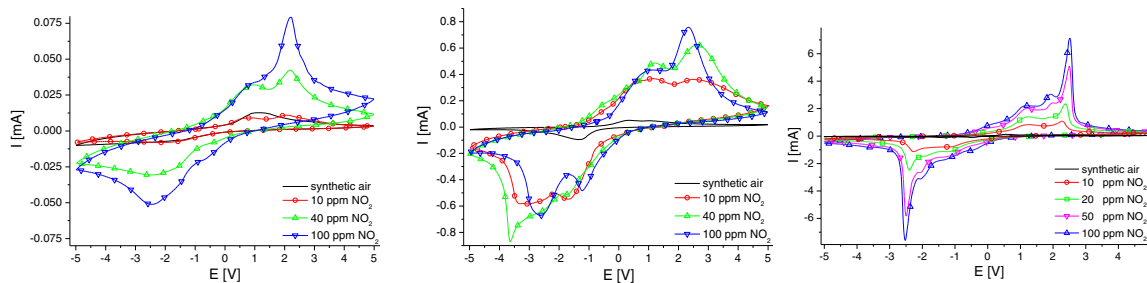


Fig. 3. Cyclic voltammograms at 300°C measured for sensor 1 (a), sensor 2 (b) and sensor 3(c)

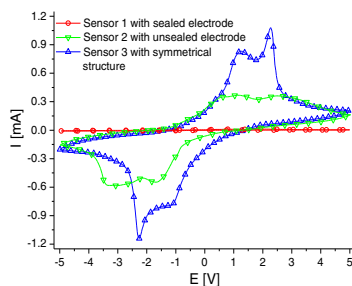


Fig. 4. Comparison between cyclic voltammograms measured for all sensors in 10 ppm NO<sub>2</sub>

### 3. Results

#### 3.1. Cyclic voltammetry

Current-voltage plots measured for all sensors in the atmosphere of the mixtures of synthetic air and nitrogen dioxide are shown in the Fig. 3. It is visible that in case of all sensors shape of obtained plots depends on the concentration of NO<sub>2</sub>. In the Fig. 4 shapes of the I-V plots measured in the atmosphere containing 10 ppm of NO<sub>2</sub> for all investigated sensors have been compared. I-V plots measured for sensor 3 (Fig. 3c) with symmetrical structure are almost symmetrical. Characteristics registered for two sensors with additional reference layers are less symmetric. For sensor 1 with sealed electrode current-voltage plots (Fig. 3a) ‘positive’ peaks are much higher than negative peaks.

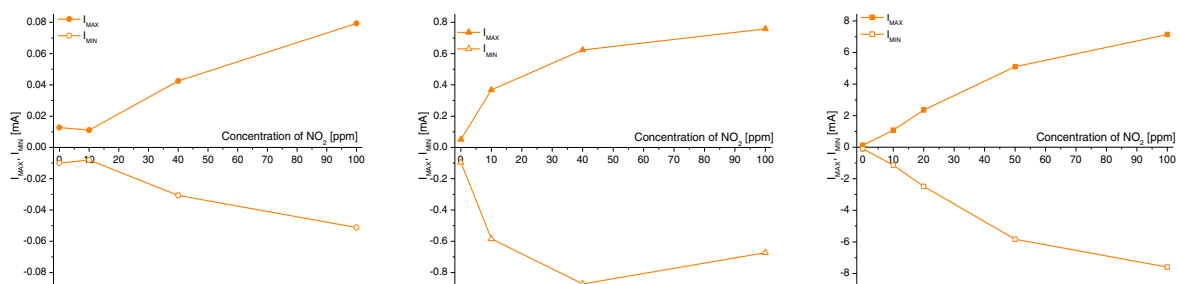


Fig. 5. Maximum current peak for different NO<sub>2</sub> concentrations sensor 1 (a), sensor 2 (b) and sensor 3 (c)

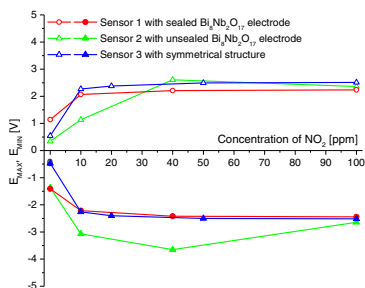


Fig. 6. Potential dependence of NO<sub>2</sub> concentration for all sensors.

### 3.2. Maximum peak current and potential dependence of NO<sub>2</sub> concentration

The maximum current peak value can be considered as a measure of the NO<sub>2</sub> concentration [4]. Maximum peak current dependences on NO<sub>2</sub> concentration are presented in the Fig. 5.  $I_{MAX}$  in the figure indicate maximum current value of 'positive' peaks and  $I_{MIN}$  maximum current value of 'negative' peak. In the most cases an increase of NO<sub>2</sub> concentration causes an increase in the current value. In the Fig. 6 values of potentials corresponding to the  $I_{MAX}$  and  $I_{MIN}$  are presented. This is the potential, in which chemical reaction occurring on the sensor electrodes is the fastest. This potential is almost independent of the NO<sub>2</sub> concentration and sensor structure. Only in case of sensor 2 values of potentials corresponding with  $I_{MIN}$  are slightly different.

Value of registered currents strongly depends on structure of the sensor. The highest values of the currents were recorded for the sensor without additional layers. This can be explained by the fact that electrochemical reaction can freely occur on both electrodes. In case of the sensor with sealed reference electrode reactions occur mainly on unsealed Pt electrode, alternately oxidation and reduction. In case of sensor 2 one of the electrodes active area is reduced by presence of Bi<sub>8</sub>Nb<sub>2</sub>O<sub>17</sub> layer.

## 4. Conclusions

In this work properties of the electrocatalytic sensors based on Nasicon and Bi<sub>8</sub>Nb<sub>2</sub>O<sub>17</sub> with different construction were investigated. Voltammetric response of these sensors in different NO<sub>2</sub> concentrations have been measured. Obtained results are compared with results of sensor with regular, symmetrical construction without any additional layers. It has been proved that presence of auxiliary layer influenced on sensor response. Based on performed experiments it is hard to judge if Bi<sub>8</sub>Nb<sub>2</sub>O<sub>17</sub> improves sensor properties. Further studies are necessary.

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