

1 Article

2 Enhanced visible-light active WO₃ thin films towards 3 air purification: effect of the synthesis conditions

4 Anna Pancielejko ¹, Marta Rzepnikowska ², Adriana Zaleska-Medynska ², Justyna Łuczak ¹,
5 Paweł Mazierski ^{2*}

6 ¹ Department of Engineering Process and Chemical Technology, Faculty of Chemistry, Gdansk University of
7 Technology, 80-233 Gdansk, Poland; anna.pancielejko@pg.edu.pl (A.P.); justyna.luczak@pg.gda.pl (J.Ł.)

8 ² Department of Environmental Technology, Faculty of Chemistry, University of Gdansk, 80-308 Gdansk,
9 Poland; marrzepn21@gmail.com (M.R.); adriana.zaleska-medynska@ug.edu.pl (A.Z.-M.)

10 * pawel.mazierski@ug.edu.pl (P.M.); Tel.: +48-58-523-52-30

11 Received: date; Accepted: date; Published: date

12 **Abstract:** Taking into consideration, our current environmental situation in the world people should face with
13 growing problem of air and water pollution. Heterogeneous photocatalysis is highly promising tool to
14 improve both, air and water quality through decomposition/mineralization of contaminations directly into
15 harmless CO₂ and H₂O under ambient conditions. In this contribution we focused on the synthesis of self-
16 assembly WO₃ thin films *via* an electrochemical approach in the aqueous electrolyte containing fluoride ions
17 towards air purifications. The effect of preparation conditions such as applied potential (10 – 50 V),
18 anodization time (15 – 120 minutes), concentration of H₂SO₄ (0.5 – 1.5 M) and NaF (0.1 – 1.0 wt.%) on the
19 morphology, photocurrent response and photocatalytic activity addressed to removal of air pollutant in the
20 presence of as-prepared WO₃ samples were thoroughly examined and presented. The results revealed the
21 growth of nanoplatelets and their gradual transformation into flower-like structure. The oxide layers and
22 platelets thickness of the WO₃ samples were found to be proportionally related with the synthesis conditions.
23 The photocatalytic ability towards air purification was evaluated by degradation of toluene from air mixture
24 using low-powered LEDs as an irradiation source ($\lambda_{max} = 415$ nm). The highest photoactivity was achieved
25 in presence of the sample which possessed well-ordered, regular shape and repeatable distribution of flower
26 buds (100% of degradation). The results have confirmed that the oxide layer thickness of the anodic WO₃
27 significantly affected the photocatalytic activity, which increased with the increasing thickness of WO₃ (to
28 1.05 μ m) and then had a down trend. The photocurrent response evidenced that well-organized sample had
29 the highest ability in photocurrent generation under UV-Vis and Vis irradiation. Finally, a possible growth
30 mechanism of WO₃ NFs was also discussed.

31 **Keywords:** WO₃ nanoflowers; electrochemical anodization; toluene degradation; visible light
32 induced photocatalysis;
33

34 1. Introduction

35 Air pollutions are the major cause of global environmental threats. They contribute to the
36 formation of respiratory diseases, infectious diseases, acute toxic effects, cancers as well as
37 reproductive disorders and allergies. The World Health Organization reports that every year over 4.2
38 million deaths is a result of exposure to ambient air pollutant[1]. Moreover, air pollutants adversely
39 affect the world, disrupting photosynthesis, transpiration and respiration. Secondly, they
40 contaminate water and soil. They increase the acidity of drinking water as well as content of heavy
41 metals such as lead, copper, zinc or cadmium in the water. Acidified water destroys plumbing
42 installations, washing away various toxic substances from it. They cause corrosion of metals and
43 building materials. Air pollutants have also negative impact on climate change. The main problem
44 which needs to be solved today it is a successful decomposition/mineralization of air contamination.
45 Therefore, an innovative and effective method of removing harmful substances from environmental
46 are sought.

47 Advanced oxidation processes (AOPs) have been extensively studied to reduce a great variety
48 of pollutants present in various environmental media[2–4]. Among them, special attention has been
49 paid for the application of heterogeneous photocatalysis in environmental applications for the
50 remediation of polluted air and water. The photocatalytic effects exerted by redox reaction are caused
51 by photoinduced of electrons (e^-) from the valence band to the conduction band leaving positive holes
52 (h^+) behind in very short time (femtoseconds) under the influence of UV-Vis irradiation. Several
53 reactive species, including formation of the high reactive hydroxyl radicals ($\bullet OH$) and superoxide
54 radicals ($O_2^{\bullet -}$), are generated through reaction of e^- and h^+ which are considered to be involved in the
55 oxidative and reductive reactions in photocatalysis. The generated reactive oxygen species react with
56 adsorbed gas and/or water pollutants leading to the degradation/mineralization them into CO_2 and
57 H_2O [5,6]. Successful photocatalytic air purification relies on the following parameters: (i) photon
58 absorption of semiconductor photocatalysts and (ii) surface properties of photocatalysts, in
59 particular, keeping clean photocatalyst surface free from the accumulation of recalcitrant
60 intermediates and products during the processes[7,8].

61 Currently, the most studied photocatalyst which holds a great potential as very effective
62 photoinduced photocatalyst is frequently used to decompose organic and inorganic contaminations
63 from water and air due to its good chemical stability, high oxidizability, non-toxicity and low-cost
64 preparation of nano- and microparticles is titanium dioxide (TiO_2). However, the main drawback of
65 it is wide band gap (3.2 eV for anatase) and consequently only small fraction of the solar irradiation
66 can be absorbed by this material ($\sim 4\%$)[9,10]. Many studies reported that sufficient doping of cations
67 or anions into TiO_2 layers to create certain states within the band gap energy may led to achieved
68 visible light driven TiO_2 photocatalysts[11]. However, those modification have a number of fails, such
69 as its thermal instability and formation of recombination centres for photoinduced charge carriers,
70 which significantly decrease its photocatalytic ability[12]. Therefore, designing and/or further
71 development of an efficient visible light active photocatalyst is particularly critical need for air
72 purification.

73 From a practical point of view, tungsten trioxide (WO_3) became an alternative photocatalytic
74 material with interesting optical, electrical and structural properties [13]. In terms of chemical
75 stability and inertness, WO_3 exhibits the same attributes as TiO_2 . In addition, it possess a lower band
76 gap energy level (between 2.4 and 2.8 eV), stable physicochemical properties and strong resilience to
77 photocorrosion effects[10,14]. Furthermore, it has been considered a great interest due to ability to (i)
78 decontaminate polluted water[15,16], (ii) detect hazardous gases[17–19] and (iii) solar energy
79 conversion[20,21].

80 Numerous studies reported different synthesis approach of WO_3 nanostructures including sol-
81 gel technique[22–24] electrochemical deposition[25], chemical vapour deposition[26,27],
82 electrochemical oxidation[28,29], magnetron sputtering[30–32], ion-beam evaporation[33,34] and
83 atomic layer deposition[35] have been described. Among them, an anodization technique attracts
84 considerable attention due to low cost and simplicity of a synthesis route. Furthermore, the possibility
85 of controlling and adjusting a growth process by tuning anodization parameters (such as applied
86 potential, synthesis time and temperature) as well as electrolyte composition makes this technique
87 more interesting for fabricating metal oxide nanostructures with controllable pore size, good
88 uniformity, and conformability over large areas[36–38]. However, several studies have reported that
89 it is still a challenge to obtain well-aligned and uniformity of anodic WO_3 nanostructures instead of
90 nanoporous[13,39–41]. First mention about preparation of porous WO_3 using galvanostatic
91 anodization of W foil in oxalic acid electrolyte, was described by Grimes *et al.*[42]. It was found that
92 the obtained nanoporous tungsten oxide exhibited more regular surface with smaller pore size in
93 compare with results described previously for other methods. Self-organized nanoporous structure
94 by anodizing W foil in electrolyte composed of 1 M sulphate acid and 0.5 wt.% sodium fluoride was
95 reported by Schmuki *et al.*[37]. The ability to control the synthesis conditions allowed them to obtain
96 desired self-ordered oxide structure exhibiting higher photocurrent efficiency than a compact oxide
97 layer. Sadek *et al.* described the growth process of WO_3 nanoplatelets during the anodization of
98 tungsten foil in a nitric acid environment at higher temperatures[43]. The combination of the

99 synthesis parameters with high temperature led the authors to obtained relatively thick films of
100 nanoplatelets with potential application in photosensitive devices. Flower-shaped tungsten oxide
101 nanostructure prepared in an acidified electrolyte solution containing fluoride ions was reported by
102 Amal *et al.*[29,44]. The resultant thin oxide film, with enhanced surface area and thickness, exhibited
103 a higher photocurrent density. What is more, the annealing temperature and crystallite of the as-
104 anodized WO₃ nanostructures was found as a critical factor e. g. in water splitting reaction[44].
105 Despite many reports describing synthesis approach of WO₃ nanostructures with electrochemical
106 method, there is still lack of knowledge about the effect of synthesis parameters (such as electrolyte
107 composition, applied potential and time period duration) on the geometrical parameters, optical and
108 photocatalytic properties with photoactivity of the WO₃ nanoflowers.

109 Motivated by this, we propose, for the first time, a novel synthesis method of the self-assembly
110 three-dimensional tungsten oxide nanoflowers (3D WO₃ NFs) thin film with improved optical and
111 photoelectrochemical properties and exhibited enhanced photocatalytic activity in the reaction of
112 toluene degradation under visible irradiation. The objectives of the present study are, therefore: (i) to
113 optimize synthesis conditions of the WO₃ NFs, (ii) to correlate morphological dimensions and
114 photoactivity and (iii) to find optimal synthesis conditions to assure that the nanostructures do not
115 decompose under illumination and its remains stable in time.

116 2. Materials and Methods

117 2.1 Materials

118 Isopropanol (p.a., POCh. S.A., Gliwice, Poland), acetone, methanol (p.a., P.P.H. STANLAB,
119 Lublin, Poland) and deionized water (DI, with conductivity of 0.05 μS) were used during sonication
120 process. WO₃ NFs were synthesis with using anodic oxidation of tungsten foil (0.127 mm, 99.9%
121 purity, Sigma Aldrich, Sigma Aldrich) in the aqueous electrolyte composed of sulphuric acid solution
122 (96%, p. a., P.P.H. STANLAB, Lublin, Poland) and sodium fluoride (p. a., P.P.H. STANLAB, Lublin,
123 Poland).

124 2.2 Synthesis of WO₃ nanoflowers

125 W foils were cut into 2 × 2 cm samples and ultrasonically cleaned in acetone, isopropanol,
126 methanol and deionised water for 10 minutes in each solvent and then dried in an air stream[45,46].
127 The as-cleaned samples were contacted with a Cu spring and pressed against an O-ring in an
128 anodization cell, which consists of a two-electrode configuration with W foil as the working electrode
129 (anode) and the platinum foil as the counter electrode (cathode). All the experiments were performed
130 at room temperature. The electrolytes containing various concentration of sodium fluoride (0.1, 0.2,
131 0.5, 0.7 and 1.0 wt.%) and sulphuric acid (0.5, 1 and 1.5 M) were used. A sets of analysis were
132 conducted for 15, 30, 45, 60, 90 and 120 minutes during anodization at the voltage range of 10 – 50 V
133 using programmable power supply (MCP M10-QS1005). The as-anodized samples were raised with
134 deionized water, dried overnight at 60°C and then annealed at 400°C for 4 h with a ramping rate of
135 4°C/min in the air environment.

136 2.3 Material characterization

137 A high resolution scanning electron microscopy (HRSEM, JEOL, JSM – 7610F) was used to
138 analysed the morphology parameters of WO₃ NFs. Cross-sectional images were obtained in which
139 the samples were scrunched and measured at the angle of 30° tilted view to determine the thickness
140 of the films. The crystal structure of the samples was determined from X-Ray diffraction patterns
141 recorded in the range of 2θ = 20 – 80°, using an X-Ray diffractometer (XRD, Rigaku MiniFlex 600)
142 with Cu Kα radiation. UV-Visible absorption spectra of the samples in the wavelength of 200 - 800 nm
143 were collected by using diffuse reflectance UV–Vis spectroscopy (Thermo Scientific) equipped with
144 an integrating sphere with baseline determined with barium sulphate as reference. The
145 photoluminescence (PL) measurements were carried out at room temperature using a
146 photoluminescence spectrometer LS-50B (PerkinElmer) equipped with Xenon discharge lamp as an

147 excitation source and a R928 photomultiplier as detector. The excitation radiation (300 nm) was
148 directed on the surface of the samples at an angle of 90°. The Raman spectra were collected with a
149 Thermo Scientific DXR Smart Raman spectrometer with a 532 nm laser as the excitation source under
150 ambient conditions.

151 2.4 Photocatalytic performance

152 The photocatalytic activity of the as-prepared WO₃ thin films was tested by the visible-light-
153 driven degradation of toluene (200 ppm) from an air mixture, used as a model contaminant. The
154 measurements were carried out in a flat stainless steel reactor of a volume of ca. 35 cm³ equipped
155 with a quartz window, two valves and a septum. The irradiation source consisted of a LED array
156 with $\lambda_{\max} = 415$ nm was located above the sample. The as-anodized foil was placed at the bottom side
157 and the reactor was closed with a quartz window. Subsequently, the gaseous mixture was passed
158 through the reactor for 1 min. Then, the valves were closed and the reactor was kept in dark for
159 30 min in order to achieve equilibrium. Before starting the irradiation, a reference toluene sample was
160 taken. The concentration was determined by using a gas chromatograph (TRACE 1300, Thermo
161 Scientific), equipped with an ionization flame detector (FID) and a Phenomenex capillary column
162 (30 mm x 25 mm, 0.5 μ m). The samples (200 μ L) were dosed with a gastight syringe each 10 min.
163 Intensity of irradiation was measured by an optical power meter and reached 15 mW/cm².

164 2.5 Photoelectrochemical activity

165 Photocurrent measurements were performed using an AutoLab PGSTAT 204 potentiationstat-
166 galvanostat (Metrohm) with the three-electrode system. Prepared samples were used as working
167 electrodes with Ag/AgCl/0.1 M KCl and Pt mesh as the reference and counter electrodes, respectively.
168 The active surface area of the electrode was 0.25 cm². Prior to the measurement, the electrolyte 0.1 M
169 Na₂SO₄ water solution, was purged with argon for 1 h. Similarly, the space above the electrolyte
170 during the measurements was purged with argon. Photocurrent measurements under UV-Vis and
171 visible irradiation were analyzed using a 250 W Xe light source with a 420 nm cut-off filter (for visible
172 light irradiation).

173 3. Results

174 To investigate the effect of the synthesis conditions and electrolyte composition, (anodization
175 potential, reaction time, concentration of H₂SO₄ and NaF) on the morphological, photoactivity and
176 photoelectrochemical properties of the anodic oxide WO₃, a series of the samples were prepared. The
177 labels of the as-prepared samples together with the synthesis conditions, morphological parameters,
178 optical band gap and efficiency of toluene degradation are presented in Table 1.

Table 1. Sample label, morphology parameters, optical band gap and efficiency of toluene degradation of the obtained samples.

Sample label	Preparation conditions (electrolyte composition and anodization parameters)		Flower diameter (nm)	Oxide layer thickness (μm)	Platelets thickness (nm)	Optical band gap (eV)	Efficiency of toluene degradation after 60 minutes of irradiation (%)
	Constant parameters	Variable parameters					
WO_10 V	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 90 min	10V	-**	0.1±0.1	11.8±0.1	2.75	36
WO_20 V	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 90 min	20V	-**	0.3±0.1	12.8±0.2	2.61	67
WO_30 V	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 90 min	30V	0.72±0.1	0.8±0.1	13.6±0.2	2.65	90
WO_40 V*	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 90 min	40V	1.21±0.1	1.05±0.1	14.9±0.3	2.52	100
WO_50 V	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 90 min	50V	0.81±0.2	1.2±0.1	16.9±0.3	2.65	73
WO_15 min	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 40V	15 min	-**	0.1±0.02	11.2±0.1	2.66	15
WO_30 min	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 40V	30 min	0.91±0.1	0.3±0.05	12.3±0.1	2.64	25
WO_45 min	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 40V	45 min	1.04±0.2	0.5±0.1	13.1±0.2	2.56	55
WO_40 V*	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 40V,	90 min	1.21±0.1	1.05±0.2	14.9±0.2	2.52	100
WO_120 min	1.0 M H ₂ SO ₄ , 0.5 wt.% NaF, 40V	120 min	1.52±0.2	1.35±0.1	15.6±0.2	2.40	78
WO_0.5 M H ₂ SO ₄	0.5 wt.% NaF, 40V, 90 mi	0.5 M H ₂ SO ₄	1.15±0.1	1.17±0.1	13.4±0.2	2.63	86
WO_40 V*	0.5 wt.% NaF, 40V, 90 min	1.0 M H ₂ SO ₄	1.21±0.1	1.05±0.1	14.9±0.2	2.52	100
WO_1.5 M H ₂ SO ₄	0.5 wt.% NaF, 40V, 90 min	1.5 M H ₂ SO ₄	1.10±0.1	1.0±0.1	15.3±0.3	2.55	90
WO_0.1 wt.% NaF	1.0 M H ₂ SO ₄ , 40V, 90 min	0.1 wt.% NaF	-**	1.0±0.01	11.3±0.1	2.79	26
WO_0.2 wt.% NaF	1.0 M H ₂ SO ₄ , 40V, 90 min	0.2 wt.% NaF	-**	1.02±0.1	12.1±0.2	2.62	61
WO_40 V*	1.0 M H ₂ SO ₄ , 40V, 90 min	0.5 wt.% NaF	1.21±0.1	1.05±0.01	14.9±0.3	2.52	100
WO_0.7 wt.% NaF	1.0 M H ₂ SO ₄ , 40V, 90 min	0.7 wt.% NaF	-**	1.1±0.01	15.9±0.3	2.55	86
WO_1.0 wt.% NaF	1.0 M H ₂ SO ₄ , 40V, 90 min	1.0 wt.% NaF	-**	1.15±0.01	17.8±0.3	2.66	74

*The WO_40 V sample was synthesised in the following conditions: 90 min, 40 V, 0.5 wt.% NaF and 1.0 M H₂SO₄ in each series to compare the influence of preparation conditions.

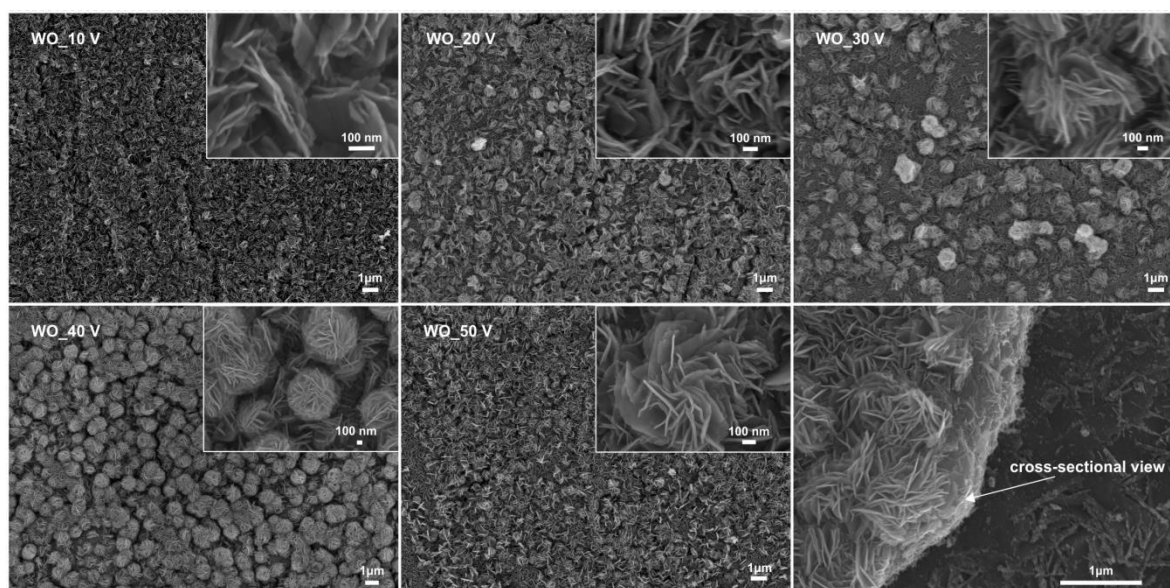
**The formation of the flower buds was not observed.



183 3.1. Morphology

184 3.1.1 Effect of the anodization potential

185 Series of samples in the range of 10 to 50 V were synthesized to determine the effect of the
 186 anodization potential on the growth of the WO₃ nanostructure. The results are displayed in Figure 1.
 187 Too low anodization potential (below 20 V) resulted in the formation of the anodic oxide layer
 188 possessing platelets nanostructure. Its irregular shape has been gradually transformed into flower-
 189 like structure when the anodization potential increased to 30 V and revealed growth of the irregular,
 190 sparsely spread flower buds with the diameter of $0.72 \pm 0.1 \text{ nm}$. Regular shape and repeatable
 191 distribution were observed for the sample anodized at 40 V with the diameter of the flowers buds
 192 equaled $1.21 \pm 0.1 \text{ }\mu\text{m}$. However, further increase to 50 V caused decrease of the flowers buds
 193 abundance (see Figure 1). Decrease in the diameter of WO₅₀ V to $0.81 \pm 0.2 \text{ nm}$ was also observed.
 194 Moreover, it was found that the nanoplatelets started to deform – were rounded and looked like
 195 developed buds. On the other hand, increasing the anodization potential resulted in an increase in
 196 the oxide layer thickness and platelets thickness from $0.1 \pm 0.1 \text{ }\mu\text{m}$ and $11.8 \pm 0.1 \text{ nm}$ for the WO₁₀ V
 197 sample to $1.2 \pm 0.1 \text{ nm}$ and $16.9 \pm 0.3 \text{ nm}$ for WO₅₀ V, respectively. Additionally, higher applied
 198 voltage resulted in sharpening the edges of the nanoplatelets.



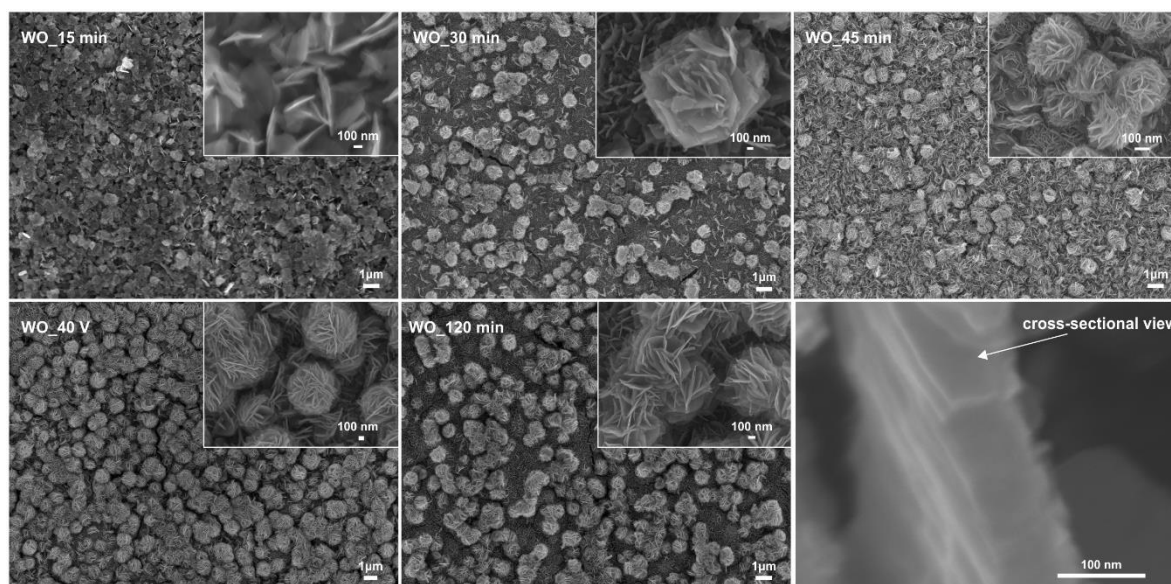
199

200 **Figure 1.** Effect of the anodization potential on the morphology of WO₃ NFs. Fluoride content
 201 0.5 wt.%, H₂SO₄ concentration 1.0 M, anodization time 90 min.

202 3.1.2 Effect of the anodization time

203 To investigate the effect of the anodization time on the growth of the nanoflowers, a series of
 204 experiments were conducted for different durations and the results are presented in Figure 2. The
 205 sample anodized at 40 V has been chosen to explore further its anodization time because it exhibited
 206 the most ordered, regular shape of flower-like structure. It established that the obtained layers
 207 consisted of irregular buds when the anodization time decreased below 90 minutes. Moreover, in
 208 shorter time fewer pores were formed and an increase in heterogeneity of the layers was observed. It
 209 could be assume that the growth of well-ordered flowers buds require soluble species which are
 210 formed by the initial anodic growth of the oxide layer, which will be further discussed in this paper.
 211 We suggest that the continuously increase in the anodization time allows to achieve steady state while
 212 the diameter of the flower buds, the thickness of the oxide layer and the platelets thickness were still
 213 improved. Diameter increased to 1150 nm, oxide layer thickness from 0.1 ± 0.02 to $1.35 \pm 0.1 \text{ }\mu\text{m}$, and

214 platelets thickness from 11.2 ± 0.1 to 15.6 ± 0.2 nm, respectively for WO_15 min and WO_120 min (see
 215 Table 1). It was found the flowers buds were evenly distributed on the layer for the anodizing time
 216 of 90 min (Figure 2). However, further extension of the anodization time had a negative effect on
 217 the WO₃ nanostructures, whereby the regular flower-like nanostructure of WO₃ was destroyed and
 218 the anodic oxide was composed of irregular, poorly distributed flower buds. Moreover, slow
 219 transformation of the flower buds again into nanoplatelets was observed. The reason might be
 220 attributed to the higher etching rate on the WO₃ surface layer with extension of the anodization time
 221 up to 120 minutes (see Fig. 2). Similar observation of the flower-like structure growth with an
 222 extension of the anodization time was also studied and published by Amal group[29].

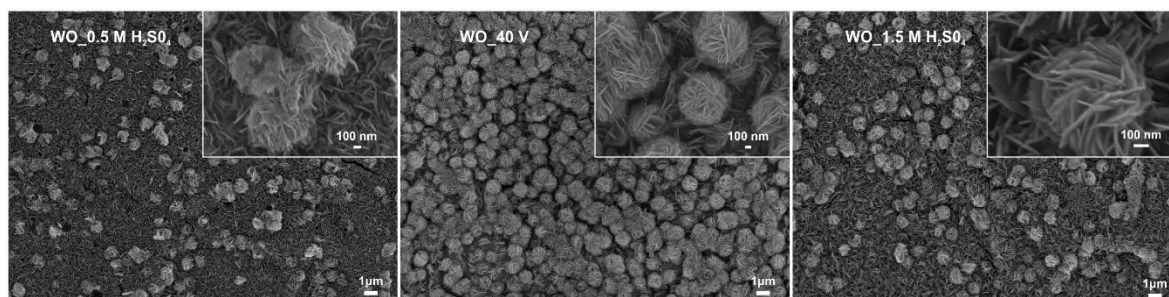


223

224 **Figure 2.** Effect of the anodization time on the morphology of the WO₃ NFs. Fluoride content 0.5 wt.%,
 225 H₂SO₄ concentration 1.0 M, anodization potential 40 V.

226 3.1.3 Effect of the sulphuric acid concentration

227 A significant impact of the sulphuric acid concentration on the homogeneity of the samples was
 228 noticed. Too low (0.5 M) or too high (1.5 M) concentration caused disorganization of the nanoplatelets
 229 forming the oxide layer. Moreover, for the both mentioned concentrations, the layers consisted of
 230 evenly distributed nanoplatelets with poorly formed irregular buds. According to the literature, for
 231 the initial oxidation step occurs, the W foil needs to be in contact with the oxidizing acid which
 232 initially oxidize into WO₂²⁺ and results in the formation of the continues oxide film[29,43]. However,
 233 we assume that too low acid concentration resulted too slow nucleation rate (WO_0.5 M H₂SO₄).
 234 Furthermore, it was found that the increase in the acidity of the environment resulted in etching of
 235 the ions WO₂²⁺ ions as it was confirmed from the Pourbaix diagram [47] and thus less distributed
 236 flower-buds were observed (see Figure 3). El-Basiouney *et al.* also concluded that the dissolution of
 237 WO₃ oxide layer in acidic medium takes place, which is consistent with the below proposed
 238 mechanism. The optimal concentration for forming regular flower buds was 1.0 M reaching the
 239 highest diameter of 1.21 ± 0.1 nm. The increase of the acid concentration had a slight influence on the
 240 oxide layer and platelets thickness, the changes from 1.17 ± 0.1 to 1.0 ± 0.1 μm and from 13.4 ± 0.2 to
 241 15.3 ± 0.3 nm, respectively for WO_0.5 M H₂SO₄ and WO_1.5 M H₂SO₄, were observed.



242

243

244

Figure 3. Effect of the sulphuric acid concentration on the morphology of WO_3 NFs. Fluoride content 0.5 wt.%, anodization potential 40 V, anodization time 90 min.

245

3.1.4 Effect of the sodium fluoride concentration

246

247

248

249

250

251

252

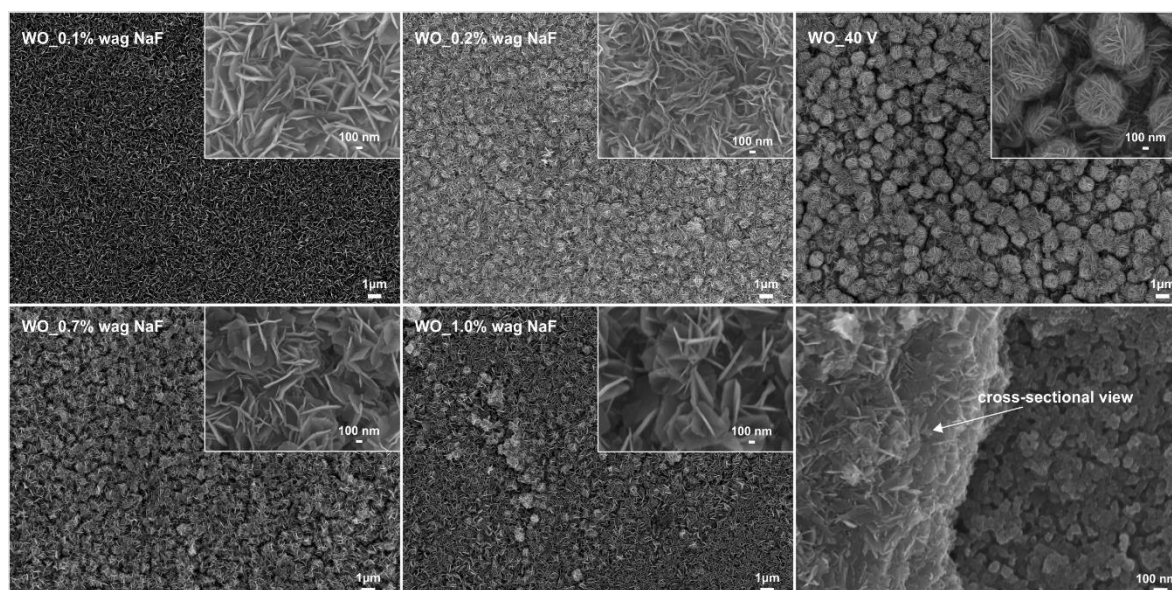
253

254

255

256

The effect of fluoride content concentration was investigated by the anodization of the samples at 40 V for 90 minutes in the electrolyte containing 0.5 M H_2SO_4 in the range of 0.1 to 1.0 wt.% NaF. The small content of F^- ions resulted in a growth of the layers consisted of the regular nanoplakelets (0.1 wt.%). The self-assembly flower buds have already occurred at 0.2 wt.% NaF content. Regularly formed buds were found for the $\text{WO}_{0.5}$ wt.% NaF sample with the diameter equalled 1.21 ± 0.1 nm. Further increased of the NaF concentration led to deformations of the flowers buds morphology to single, packed nanoplakelets. Moreover, the results revealed that with the increasing amount of NaF concentration, increase of sharpened edges of the nanoplakelets was observed. The oxide layer and platelets thickness increased slightly from 1.0 ± 0.01 to 1.15 ± 0.01 μm and from 11.3 ± 0.1 to 17.8 ± 0.3 nm, respectively for $\text{WO}_{0.1}$ wt.% NaF and $\text{WO}_{1.0}$ wt.% NaF, implying that the NaF concentration was mainly responsible for the flower buds formation.



257

258

259

Figure 4. Effect of the sodium fluoride concentration on the morphology of the WO_3 NFs. H_2SO_4 concentration 1.0 M, anodization potential 40 V, anodization time 90 min.

260

261

262

263

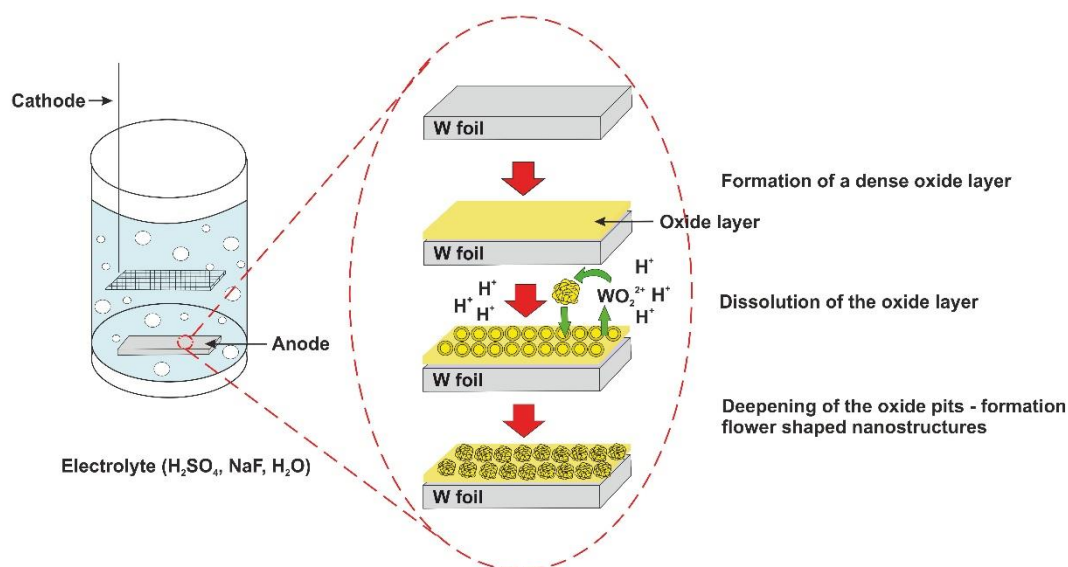
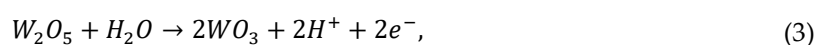
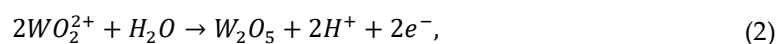
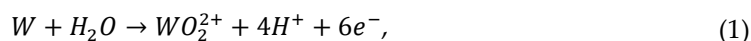
264

265

266

267

Similarly to the formation of the anodized titanium dioxide nanotubes films (TiO_2 NTs) in the electrolyte containing fluorine ions [48,49], the WO_3 NFs films were formed as a result of field-assisted oxide growth and localized chemical dissolution. Based on the above results and literature reports [28,29,50], we propose the following mechanism of the WO_3 NFs thin film growth process: (i) formation of a dense oxide layer on W foil, (ii) activation of the barrier oxide layer by fluoride ions resulting in chemical dissolution of the oxide layer and (iii) deepening of the oxide pits that in time branch out formation flower shaped nanostructures (Scheme 1). The proposed schematic illustration of the anodic growth of oxide layer is presented on Scheme 1 according to the following equations:



268

269

Scheme 1 Schematic illustration of the WO₃ NFs thin film growth process.

270

271

272

273

274

275

276

277

278

279

280

281

282

283

284

Reaction 1 describes the electrochemical dissolution of tungsten foil and formation of an oxide layer of WO₂²⁺ on the surface. Subsequently, the WO₂²⁺ ions are attracted by water molecules and form an intermediate W₂O₅ oxide. Since the WO₂²⁺ and intermediate W₂O₅ are decomposed in an aqueous environment, the oxide layer growth. Further oxide growth is controlled by the field enhanced ion transport through the growing oxide. This process is self-limiting under a constant applied voltage, as the field within the oxide layer is progressively reduced by the increasing oxide thickness, thus resulting in the growth of a compact WO₃ film with finite thickness. During the initial step, disordered pits are formed and a nanoporous structure is subsequently developed by the chemical dissolution of the oxide layer or the direct complexation of WO₂²⁺ at the oxide electrolyte interface and form soluble fluoride complexes[49]. The anodic growth of the WO₃ oxide layer became a competition between its and the chemical dissolution of the tungsten oxide layer in a fluoride containing electrolyte solution[51,52]. During the anodization process, the constant growth and chemical dissolution of the tungsten oxide layer occurs simultaneously and a steady state is established when the growth rate at the metal oxide interface occurs at the same as the dissolution rate of oxide film at the outer interface[13,28].

285

286

287

288

289

290

291

292

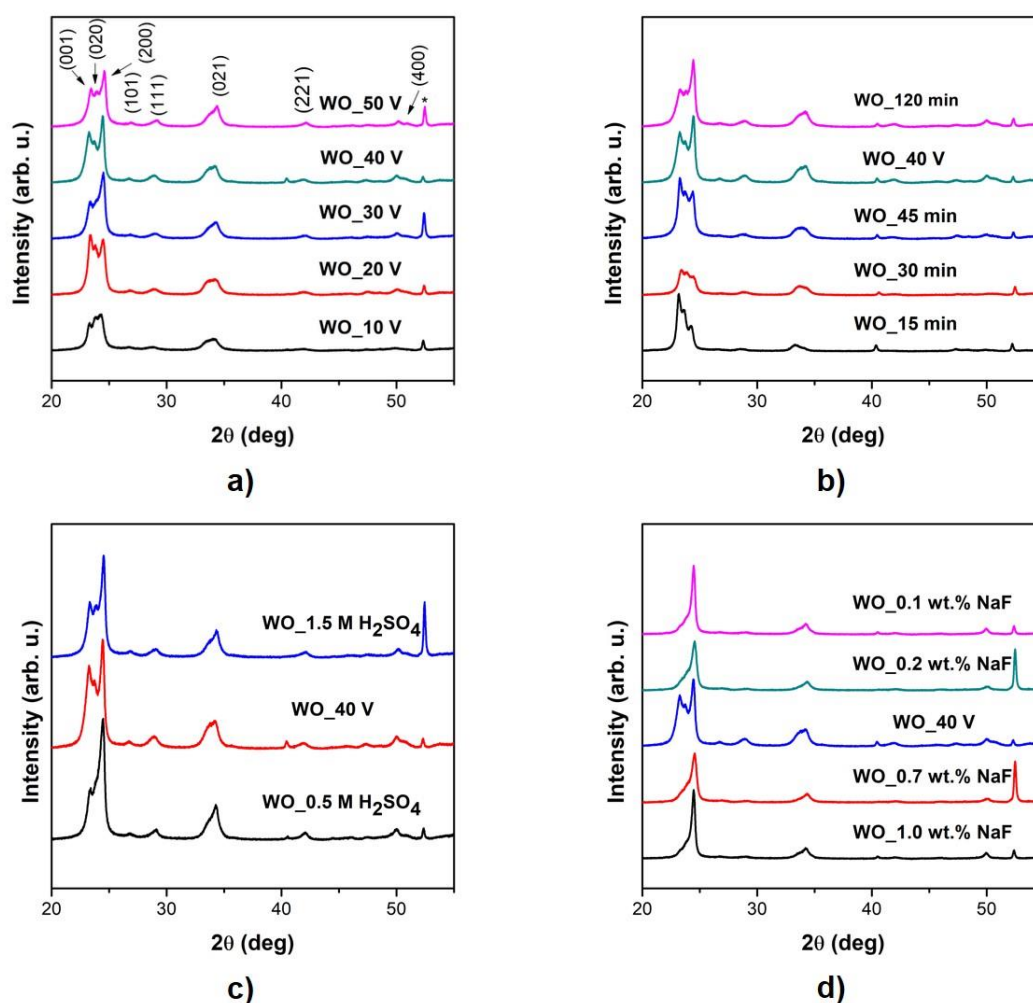
As mentioned above, the anodic growth of the compact oxide on the metal surface and the formation of pores is governed by chemical dissolution of the formed oxide layers induced by fluoride ions from the electrolyte solution and formation of a soluble fluoride complex. The key point is to find an optimize amount of the fluoride content, needed to form a porous structure, allows for the successful formation of the flower-like structure. Too low or too high concentration revealed presence of the nanoplatelets instead of nanoflowers (see Figure 4) indicating that the presence of fluoride ions is essential to generate soluble ions (WO₂²⁺) and form the flower-like structure. Moreover, an increase in the NaF concentration (to 1.0 wt.%) resulted in slowly etching of the oxide

293 layer and exhibited poorly organized flowers buds co-existence together with nanoplatelets (see
294 Figure 2). Lai [53] also investigated the influence of the fluoride content on the growth of anodic WO_3
295 nanotubular structure and realised that optimization of fluoride ions played crucial role in controlling
296 formation the chemical dissolution reaction on the interface of W/WO_3 and thus growth of the
297 nanotubes. It should be remarked that dissolution of WO_3 occurs over the entire W foil, thus with
298 extending oxidation time we observed gradual transformation of the nanoplatelets into flower-like
299 structure after 90 minutes of the anodization process. Further time extension resulted in an opposite
300 effects (Figure 2). Moreover, literature survey reports that the dissolution of WO_3 in acid medium
301 takes place *via* the formation of WO_2^{2+} species[54]. Therefore, it might lead to the precipitation of
302 primary formed WO_2^{2+} and allows for thickening of the WO_3 nanostructured film. However, we
303 observed that the increase of the acid concentration resulted in the formation of slightly thinner oxide
304 layer (see Table 1). Therefore, we assume that the dissolution of the oxide layer could be driven by
305 the instability of WO_2^{2+} in the presence of aqueous electrolyte solution at room temperature and
306 presence of high anodization potential.

307 3.2 Crystallographic structure

308 Figure 5 displays the X-ray patterns of the self-assembled WO_3 NFs film. The reflection patterns
309 of the WO_3 could be indexed to a triclinic phase. The WO_3 phase was represented by the peaks (001),
310 (020), (200), (101), (111), (021), (221) and (400) crystal planes at 2θ located around 23.3° , 23.8° , 24.4° ,
311 26.8° , 28.9° , 33.8° , 41.4° and 50.5° . The peaks indexed to W foils were found at 2θ of approximately
312 52.5° (denoted in Figure 5 “*”), 58.5° and 73.4° . The XRD patterns of the samples anodized for 90
313 minutes in the potential range of 10 – 50 V are given in Figure 5a. As noted, the potential of 10 V was
314 too low for successful formation of the oxide layers thus it was hard to establish any of the peaks in
315 range of 23° – 25° . It was found that with increasing anodization potential the intensity of the peak
316 (200), characteristic for triclinic phase of WO_3 , increased. However, further rise to 50 V resulted in the
317 intensity decrease. As the anodization time increased we observed that the intensity of the (001) peak
318 decrease whereas the (200) one started to intensively arise and reached its highest intensity with an
319 oxidation time of 120 minutes (see Figure 5b). As the sulphuric acid concentration increased, the
320 intensity of the (200) peak started to decrease whereas the intensity of peaks indexed to (001) and
321 (020) became higher (Figure 5c). An optimum concentration which allowed for the growth of the
322 oxide layer was 1.0 M. Interestingly, concentration of fluoride content strongly affected on the peaks
323 height indexed to WO_3 (Figure 5d). Only the $\text{WO}_{0.5}$ wt.% NaF sample revealed the presence of the
324 peaks (001), (020), (200), (101) and (111), whereas the rest of the samples in this series possessed the
325 intensive peak indexed to (200). Reduction of the (001) peak and thus enhancement of the (200) peak
326 intensity in all the samples could be ascribed to the improvement of the nanostructured layer and
327 compact oxide layers obtained at different conditions. The WO_3 NFs samples possessed analogous
328 crystallinity, while different synthesis conditions resulted in changes of the refined lattice parameters
329 a , b , and c as well as unit cell volume, which are gathered in Table S1.





330

331

332

Figure 5. XRD patterns of the WO_3 NFs samples differing in (a) applied potential, (b) anodization time, (c) H_2SO_4 and (d) NaF concentration.

333

334

335

336

337

338

339

340

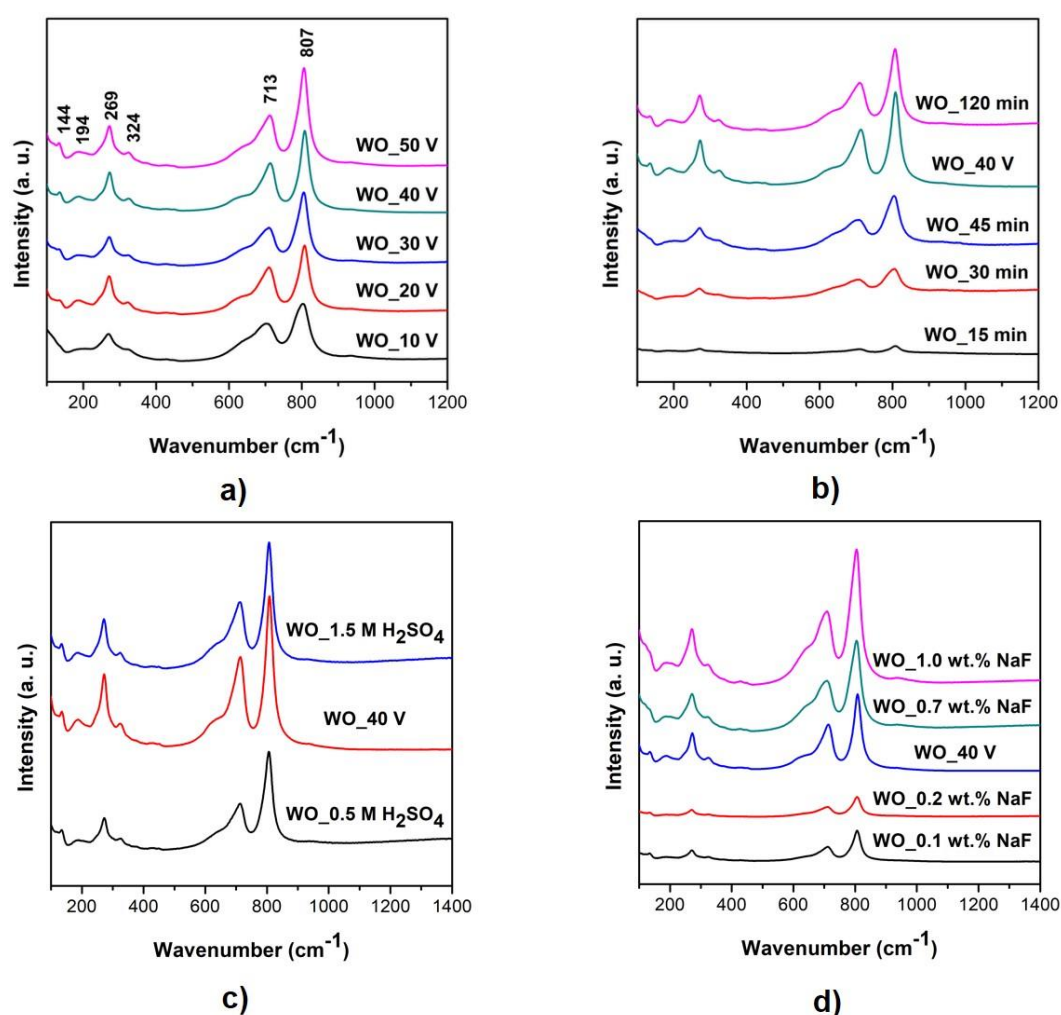
341

342

343

344

The Raman spectra of the samples are displayed in Figure 6. As it can be seen, we can determine three frequency regions. The first one appeared at lower frequencies (below 200 cm^{-1}) in regards to the relative translational or rotational motions of the WO_6 octahedral. The second region occurs between wave_numbers of $200 - 400\text{ cm}^{-1}$ indicating the O-W-O bending mode. The last region with peaks located at around $600 - 900\text{ cm}^{-1}$ is indexed to the O-W-O stretching modes[55]. The Raman spectra confirmed the WO_3 triclinic phase due to the presence of the characteristic peaks at around $144, 194, 269, 324, 713$ and 807 cm^{-1} . No Raman signal corresponding to the tungsten was observed. The bands of 269 and 324 cm^{-1} can be attributed to the δ bending (O-W-O) and ν (W-O-W) vibrations modes of the bridging oxygen[56]. The strongest peaks observed at 713 and 807 cm^{-1} can be assigned to the stretching modes arising from O-W-O[23,56]. All the samples exhibited the same features but as the preparation conditions changed, an intensity of the peaks starts to arising (especially at around 269 and 807 cm^{-1}) indicated the formation of the oxide layer (see Figure 6 a-d).



345

346

347

Figure 6. Raman spectra of the WO_3 NFs samples differing in (a) anodization potential, (b) anodization time (c) H_2SO_4 and (d) NaF concentration.

348

3.3. Optical and photoluminescence properties

349

350

351

352

353

354

355

356

357

358

359

360

361

362

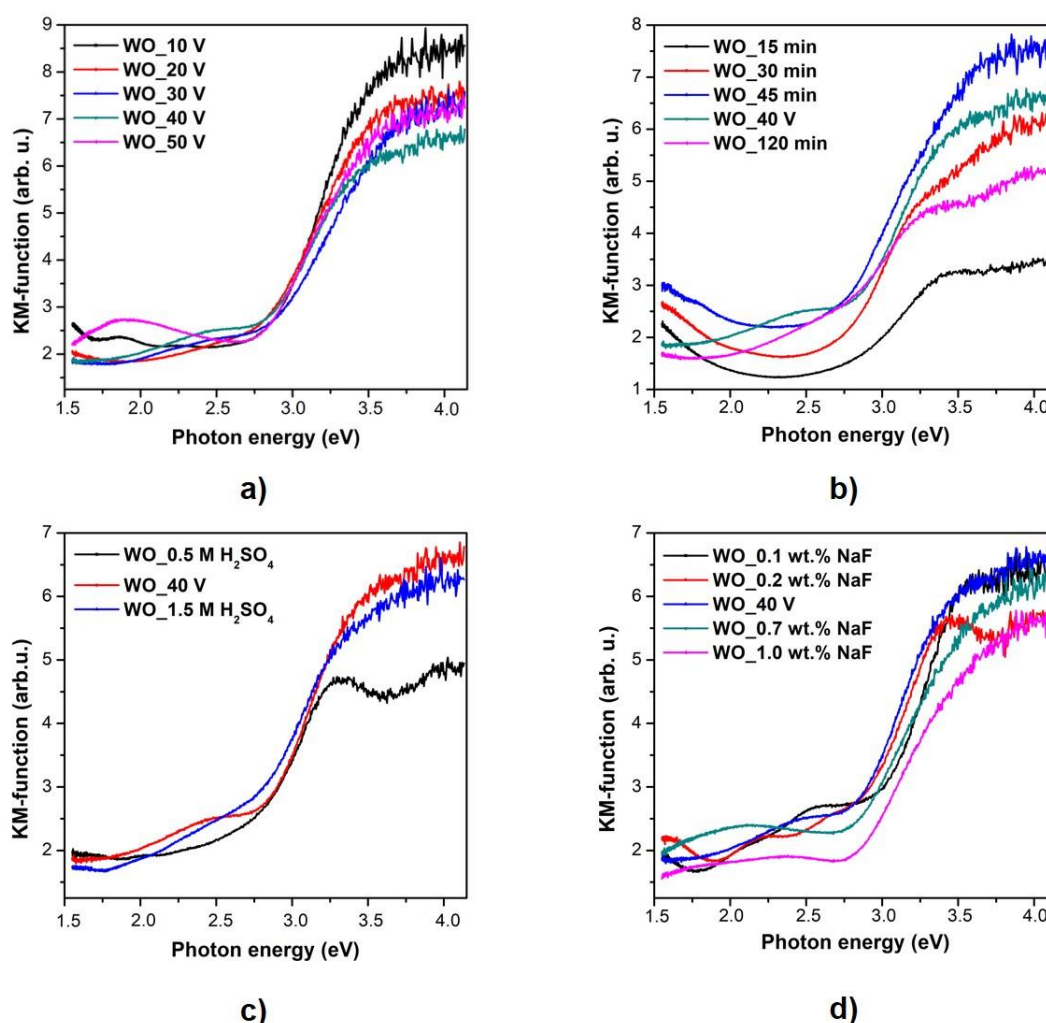
363

The change in the optical properties of the as-prepared WO_3 films by the photoabsorption studies (Figure S1) and corresponding optical band gap energy as presented in Figure 7 were investigated. The absorption edge of the series differing in the applied potential was located approximately at 470 nm for all samples (Figure S1a). As the anodization time increases from 15 to 120 minutes the absorption edge shifted from 460 to 540 nm (see Figure S1b), respectively. Higher acid concentration practically did not influence on the absorption edge which was around 475 nm (Figure S1c). For the series differing in fluoride ions concentration, the absorption edge was around 450 nm for all samples except the one obtained with the lowest amount of NaF (0.1 wt.%) – approximately 430 nm (Figure S1d). An absorption shift in the 450–800 nm range was observed for all samples, confirming changes in the structure of the obtained samples. Raised absorbance values at higher wavelengths could be attributed to the presence of oxygen vacancies on the surface film. Furthermore, it was found that as the anodization time increased the photoabsorbance values decreased due to the decrease in amount of suboxides and oxygen vacancies. A similar observation was described by Amal et al. who analysed the influence of duration time period on the flower-shaped WO_3 growth[29].

364 The band gap of WO₃ NFs films can be determined by considering the indirect transition
 365 between 2p electrons from the valence band (VB) of oxygen and the 5d the conduction band (CB) of
 366 tungsten based on the Tauc's plot according to the following equation[57,58]:

$$\alpha h\nu = A(h\nu - E_g)^n, \quad (4)$$

367 where α , ν , A and E_g are absorption coefficient, light frequency, proportionality constant, and
 368 band gap, respectively. The band gap energies of the WO₃ NFs samples were calculated with equation
 369 4 and the data were collected in Table 1 and displayed on Figure 7. The values were in accordance
 370 with the literature, where the band gap values for the flower-shaped WO₃ film was around 2.5 -
 371 2.75 eV[37,40].

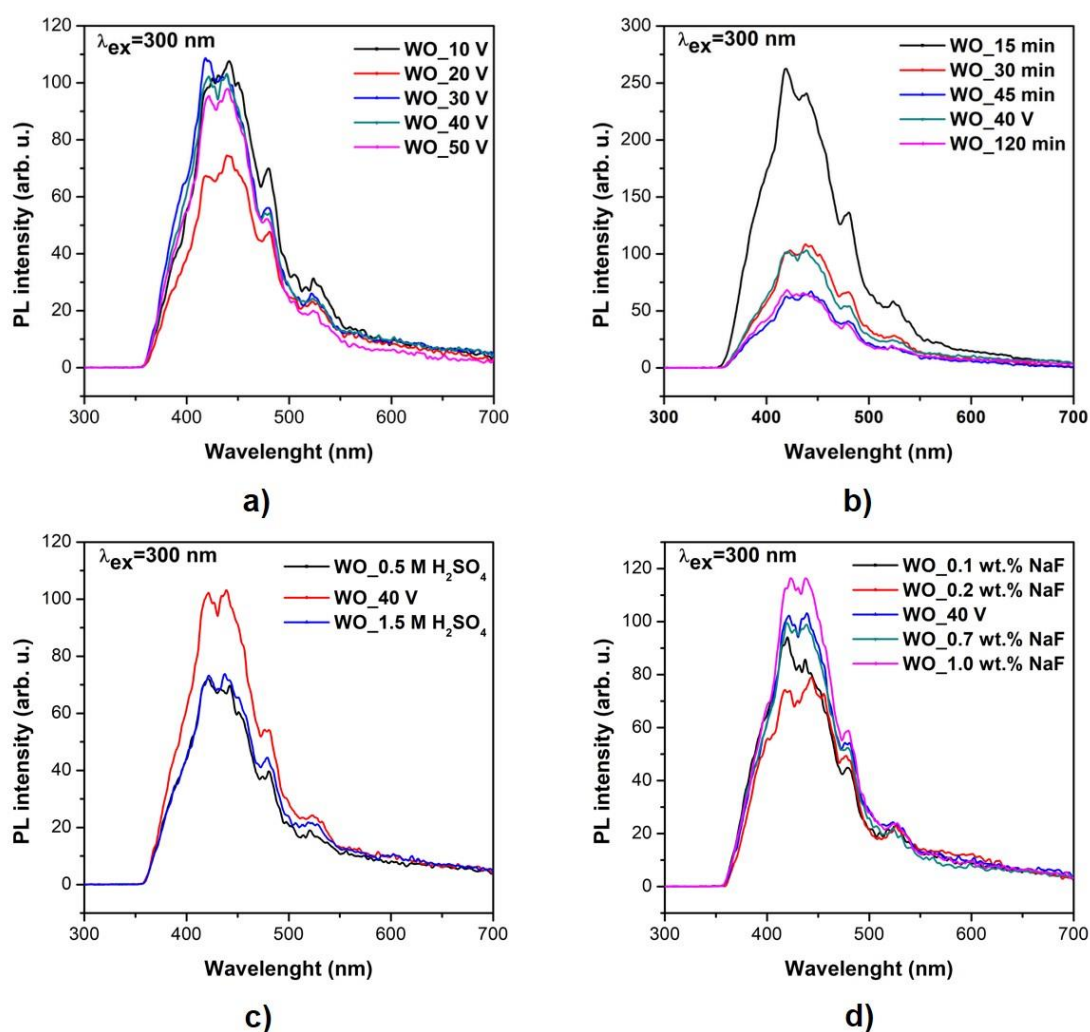


372
 373 **Figure 7.** Optical band gap spectra of the WO₃ NFs samples differing in (a) anodization potential,
 374 (b) anodization time (c) H₂SO₄ and (d) NaF concentration.

375 Photoluminescence spectra for the as-obtained WO₃ NFs samples in four different series are
 376 shown in Figure 8 a-d. Typically for the oxygen metal based semiconductors, the PL spectra are
 377 composed of UV emission and a visible emission band attributed to the surface defects. All the
 378 samples exhibited the same PL features with different intensity depending on the preparation
 379 conditions. The emission was positioned at around 420, 438, 481 and 527 nm. The values at around
 380 420, 438 and 481 nm might be attributed to the presence of intrinsic defects such as oxygen vacancies
 381 giving rise to donor states located below the CB. The emission located at wavelength of 527 nm



382 indicated the possibility of band recombination (intrinsic states rather than surface states). The values
 383 are consistent with those reported for the others semiconductors, such as TiO₂ and ZnO[59,60].

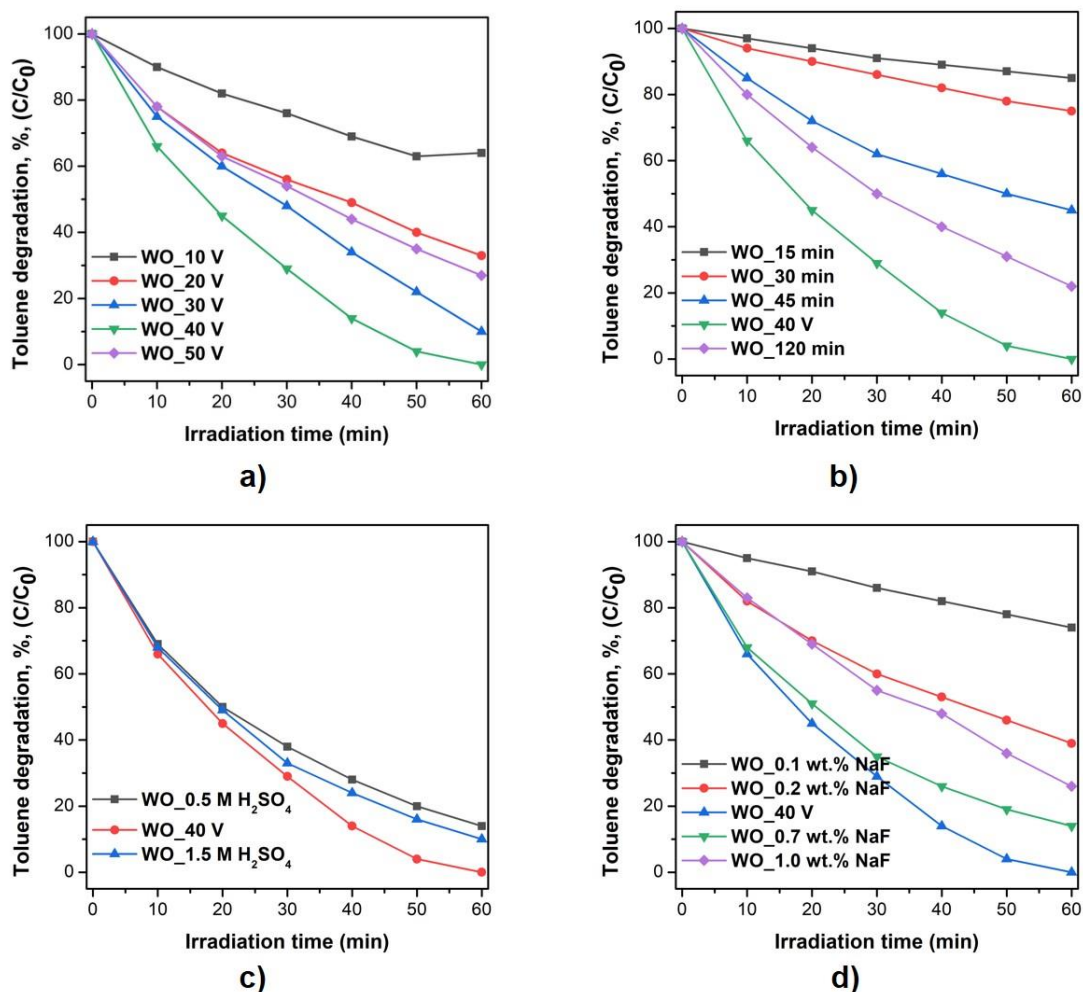


384

385 **Figure 8** Photoluminescence spectra of the WO₃ NFs samples differing in (a) anodization potential,
 386 (b) anodization time (c) H₂SO₄ and (d) NaF concentration.

387 3.4 Photocatalytic activity

388 The potential environmental applications of the as-prepared WO₃ samples were investigated in
 389 a model reaction of toluene degradation from the air mixture to simple degradation products like
 390 CO₂ and H₂O. This approach was employed to analyse the effect of the preparation conditions,
 391 anodization potential and time, as well as electrolyte composition, concentration of sulphuric acid
 392 and sodium fluoride, on the WO₃ NFs photoactivity. The obtained results are displayed in Figure 9
 393 and Table 1. It was found that samples preparation route by changing applied potential, anodization
 394 time and fluoride content significantly influenced the photocatalytic activity. In the case of different
 395 H₂SO₄ concentration, the samples WO_0.5 M H₂SO₄, WO_1.0 M H₂SO₄ and WO_1.5 M H₂SO₄ had
 396 comparable efficiency, 86%, 100% and 90%, respectively. As shown in Figure 9 a-d the highest
 397 photocatalytic activity, reaching 100% of toluene removal after 60 minutes of irradiation, achieved
 398 the sample prepared in the following conditions: anodization potential 40 V, anodization time
 399 90 min, fluoride content 0.5 wt.% and H₂SO₄ concentration 0.5 M. Besides, further increase of the each
 400 parameter resulted in an opposite effect, decrease of photoactivity was observed for WO_50 V,
 401 WO_120 min, WO_1.5 M H₂SO₄ and WO_0.1 wt.% NaF, 73%, 78%, 90% and 74%, respectively.



402

403

404

405

Figure 9 Photoactivity of the WO₃ NFs samples in gas phase degradation of toluene under Vis light irradiation ($\lambda_{\max} = 415$ nm) differing in (a) anodization potential, (b) anodization time (c) H₂SO₄ and (d) NaF concentration.

406

407

408

409

410

411

412

413

414

415

416

417

418

419

420

421

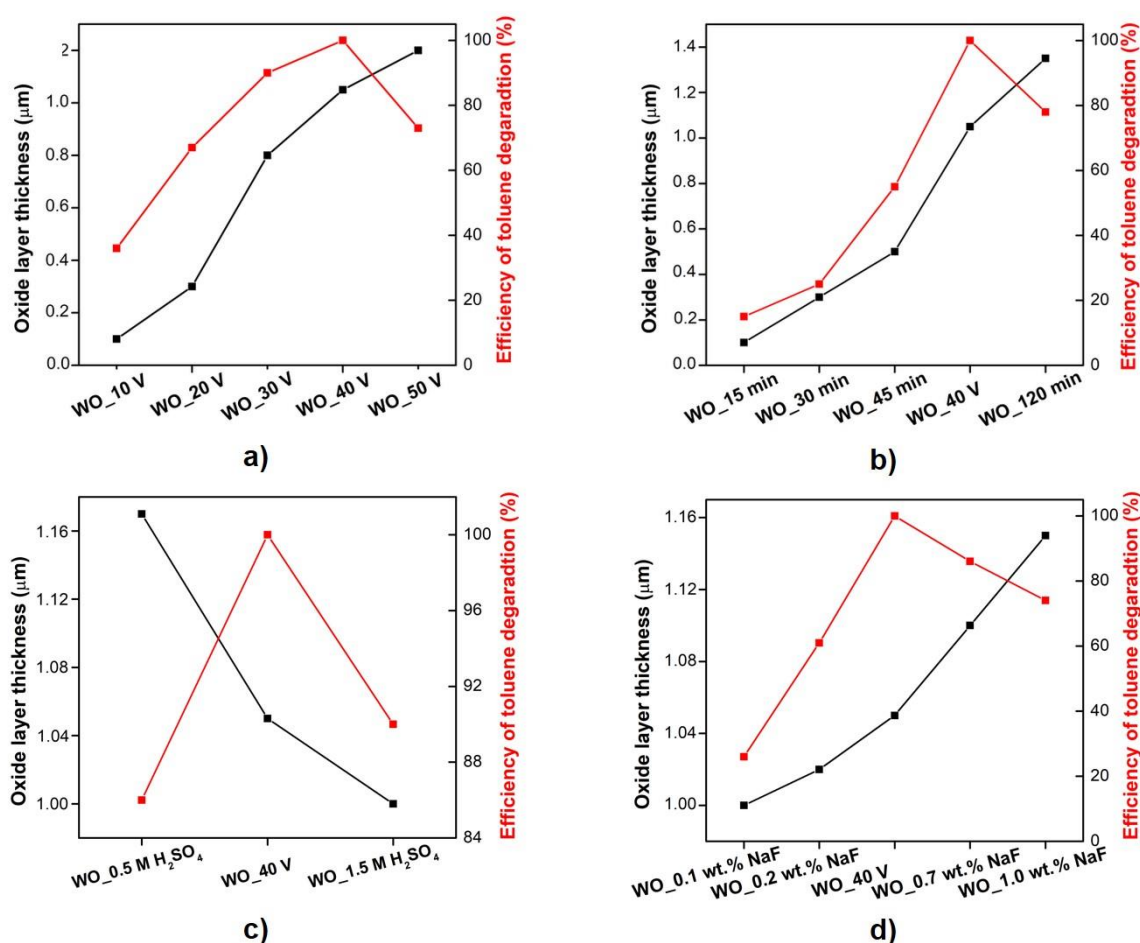
422

423

In line with the other photocatalysts, the photoactivity of WO₃ depends on the lifetime of photogenerated charge carriers. Surface electron-hole recombination is extremely high with most of charge carriers recombining on the photocatalyst surface before the redox reactions and this recombination rate must be reduced in order to improve the photocatalytic activity. It seems that a simple toluene degradation route in presence of the WO₃ samples relies on the attack of the OH• radical to the methyl group of toluene. The key parameter that influence the improvement of the photocatalytic activity towards air pollutants degradation is morphology control. According to the literature, it was found that the flower-like three-dimensional (3D) structure photocatalysts possesses (i) larger surface area and (ii) plenty of mesopores with ordered open pore frameworks, into which photocatalyst may effectively harvest visible light due to multiple scattering[61,62]. Moreover, the large surface of flower-like structure could decrease the recombination efficiency of the photoexcited carries and favours their transfer to the surface to react with organic pollutants. When comparing the photoactivity of the as-prepared WO₃ samples with their morphological parameters, it was obvious that well-ordered, regular shape of nanoflowers array films were more efficient than the random occurred flowers buds or nanoplatelets (Table 1, Figure. 10). It could be ascribed to the more effective separation for the photogenerated electron-hole pairs and the larger surface area of the nanoflowers structure[63]. As apparent from the above discussed experimental data, the highest photocatalytic activity in the reaction of toluene decomposition (100%) after 60 min of LED irradiation ($\lambda_{\max} =$



424 415 nm) was reached for the WO₄₀ V sample prepared in the following conditions: 40 V, 90 min,
 425 1.0 M H₂SO₄ and 0.5 wt.% NaF. This sample also possessed the most uniform distribution and regular
 426 shape of the flowers buds with the diameter equalled to $1.21 \pm 0.1 \mu\text{m}$. We noticed that the efficiency
 427 of toluene degradation increased with increasing thickness of WO₃ array film, and then had a down
 428 trend (Fig. 10 a-d). The experimental results indicated that the 1.05 μm thick WO₃ film appeared with
 429 a maximum photodegradation efficiency of the toluene removal. Previous literature suggests that if
 430 the metal oxide thin film is thicker than the depth of light penetration, the bottom film absorbs only
 431 few incident photons and serves as an inter support and resulted in the decrease of the photocatalytic
 432 activity for a thick film[64].



433

434

435

Figure 10. Influence of the oxide layer thickness on the efficiency of toluene degradation in the series differing in (a) applied potential, (b) anodization time, (c) H₂SO₄ and (d) NaF concentration.

436

3.5 Photoelectrochemical activity

437

438

439

440

441

442

443

444

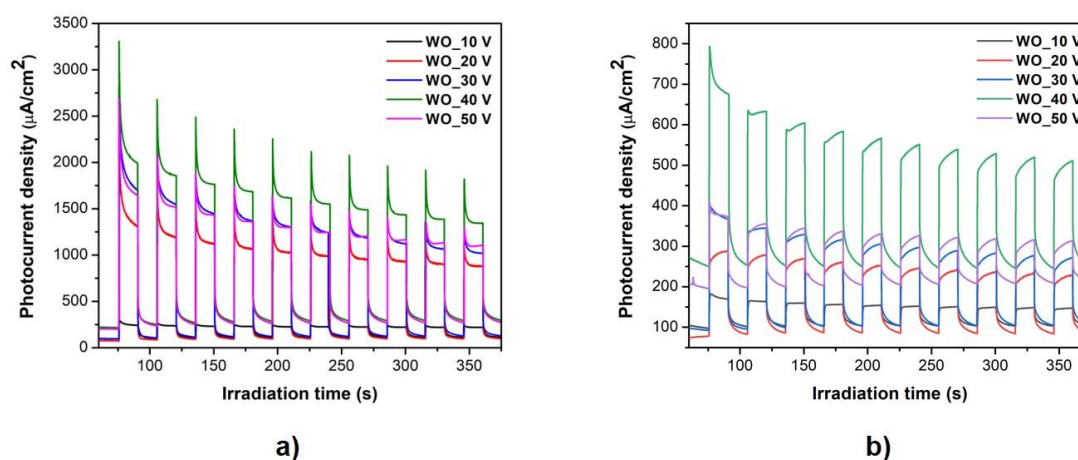
445

446

447

Photoelectrochemical response of the samples prepared at different applied potential was evaluated as photocurrent measurements in light off/on cycles with 1.5 V applied voltage. Photocurrent was registered under both UV-Vis (Figure 11a) and Vis irradiation (Figure 11b). No significant decay of the photocurrent was observed during the photoelectrochemical measurements under both sources of irradiation indicating good stability of obtained photoelectrodes. The photocurrent under UV-Vis light was approximately 5 times higher in comparison with visible light irradiation. In addition, the same tendency in sample photoelectroactivity was observed for both types of light sources, namely as the anodizing voltage increased to 40 V, the measured photocurrent increased and then decreased (for a sample obtained at 50 V). This could be related to the high ordering of the platelet structure and the improvement of the nanostructured layer thickness with larger surface area which was achieving with increasing anodization voltage (up to 40 V). Prepared

448 samples, especially the most photoactive sample (WO_40 V), exhibited the large surface of flower-
 449 like structure which increased the interface between oxide layer and electrolyte, facilitating incident
 450 photons absorbance and the charge transfer for the separation of photoexcited hole – electrons pairs.



451

452 **Figure 11.** Photocurrent response of selected WO₃ samples under a) UV-Vis and (b) Vis irradiation
 453 ($\lambda > 420$ nm) at 1.5 V in 0.1 M Na₂SO₄. The switch on/off period is 30 s.

454 5. Conclusions

455 In summary, we have described a simple and environmental friendly one-step anodization
 456 synthesis method of the WO₃ NFs thin films in the aqueous solution containing fluoride ions. The
 457 effect of synthesis conditions, applied potential and anodization time, as well as electrolyte
 458 composition, concentration of H₂SO₄ and NaF contents, on the morphology, photocatalytic activity
 459 and photoelectrochemistry properties of the WO₃ photocatalysts was presented. The results revealed
 460 the growth of the nanoplatelets which were slowly transformed into the nanoflowers structure
 461 depending on the synthesis parameters. The optimal conditions allowing the synthesis of the well-
 462 organize, regular shaped flower buds with the diameter of 1.21 μm were: anodization potential of
 463 40 V, anodization time 90 minutes, the electrolyte containing 1.0 M H₂SO₄ and 0.5 wt.% NaF. The
 464 photocatalytic activity of the as-prepared WO₃ samples towards decomposition of air contaminants
 465 was investigated in the model reaction of toluene degradation. Moreover, we found that the
 466 morphology control was the key avenue to improve the photocatalytic activity. Hence, the samples
 467 which consisted of the irregular shaped nanoplatelets exhibited lower photocatalytic activity than
 468 those with the flower-like structure. The highest photoactivity in toluene degradation and in
 469 photocurrent generation under UV-Vis and Vis irradiation achieved the sample with the most
 470 ordered and regular flower buds, WO₄₀ V. Moreover, we noticed that the thickness of the oxide
 471 layer was directly related with the photocatalytic efficiency. The thicker WO₃ layer was, the higher
 472 decline in the photoactivity was observed because of difficulties of light penetration into deeper
 473 layers. Although, the highest photoactivity exhibited the sample with thick of 1.05 μm (100%,
 474 WO₄₀ V) and further increase of the oxide layer thickness resulted in an opposite effect. These
 475 finding suggest that self-assembly, regular shape flower-like WO₃ thin film activated with low-
 476 powered LEDs as an irradiation source ($\lambda_{\text{max}} = 415$ nm) could be a promising material for air
 477 purification. The synthesis of visible light active and stable photocatalyst can boost the technology of
 478 air purification since they exploit a renewable energy source and allow to avoid an extra operational
 479 costs and other limitations associated with artificial illumination.

480 **Supplementary Materials:** Figure S1: Photoabsorption spectra of the WO₃ NFs samples from four series each
 481 with a different (a) applied potential, (b) anodization time, (c) H₂SO₄ and (d) NaF concentration., Table S1: Lattice
 482 parameters of the WO₃ NFs samples.

483 **Author Contributions:** Conceptualization, P.M.; investigation, A.P., M.R., A.Z.-M., J.L., P.M.; supervision, J.L.,
484 A.Z.-M. and P.M.; writing (review and editing) A.P., A.Z.-M., J.L., P.M.

485 **Conflicts of Interest:** The authors declare no conflict of interest.

486 References

- 487 1. who.int Available online: https://www.who.int/gho/phe/outdoor_air_pollution/burden/en/ (accessed
488 on Jul 16, 2020).
- 489 2. Ghime, D.; Ghosh, P. Advanced Oxidation Processes: A Powerful Treatment Option for the Removal of
490 Recalcitrant Organic Compounds. *Adv. Oxid. Process. - Appl. Trends, Prospect.* **2020**, 1–12,
491 doi:10.5772/intechopen.90192.
- 492 3. Dewil, R.; Mantzavinos, D.; Poullos, I.; Rodrigo, M.A. New perspectives for Advanced Oxidation
493 Processes. *J. Environ. Manage.* **2017**, *195*, 93–99, doi:10.1016/j.jenvman.2017.04.010.
- 494 4. Miklos, D.B.; Remy, C.; Jekel, M.; Linden, K.G.; Drewes, J.E.; Hübner, U. Evaluation of advanced
495 oxidation processes for water and wastewater treatment – A critical review. *Water Res.* **2018**, *139*, 118–
496 131, doi:10.1016/j.watres.2018.03.042.
- 497 5. Hoffmann, M.R.; Martin, S.T.; Choi, W.; Bahnemann, D.W. Environmental Applications of
498 Semiconductor Photocatalysis. *Chem. Rev.* **1995**, 69–96, doi:doi.org/10.1021/cr00033a004.
- 499 6. Maira, A.J.; Yeung, K.L.; Soria, J.; Coronado, J.M.; Belver, C.; Lee, C.Y.; Augugliaro, V. Gas-phase photo-
500 oxidation of toluene using nanometer-size TiO₂ catalysts. *Appl. Catal. B Environ.* **2001**, *29*, 327–336,
501 doi:10.1016/S0926-3373(00)00211-3.
- 502 7. Weon, S.; He, F.; Choi, W. Status and challenges in photocatalytic nanotechnology for cleaning air
503 polluted with volatile organic compounds: Visible light utilization and catalyst deactivation. *Environ.*
504 *Sci. Nano* **2019**, *6*, 3185–3214, doi:10.1039/c9en00891h.
- 505 8. Venkata Laxma Reddy, P.; Kim, K.H.; Kim, Y.H. A review of photocatalytic treatment for various air
506 pollutants. *Asian J. Atmos. Environ.* **2011**, *5*, 181–188, doi:10.5572/ajae.2011.5.3.181.
- 507 9. Chen, X.; Mao, S.S. Titanium dioxide nanomaterials: Synthesis, properties, modifications and
508 applications. *Chem. Rev.* **2007**, *107*, 2891–2959, doi:10.1021/cr0500535.
- 509 10. Kubacka, A.; Fernández-García, M.; Colón, G. Advanced nanoarchitectures for solar photocatalytic
510 applications. *Chem. Rev.* **2012**, *112*, 1555–1614, doi:10.1021/cr100454n.
- 511 11. Nah, Y.C.; Paramasivam, I.; Schmuki, P. Doped TiO₂ and TiO₂ nanotubes: Synthesis and applications.
512 *ChemPhysChem* **2010**, *11*, 2698–2713, doi:10.1002/cphc.201000276.
- 513 12. Wang, Y.; Cheng, H.; Hao, Y.; Ma, J.; Li, W.; Cai, S. Photoelectrochemical properties of metal-ion-doped
514 TiO₂ nanocrystalline electrodes. *Thin Solid Films* **1999**, *349*, 120–125, doi:10.1016/S0040-6090(99)00239-4.
- 515 13. Guo, Y.; Quan, X.; Lu, N.; Zhao, H.; Chen, S. High photocatalytic capability of self-assembled
516 nanoporous WO₃ with preferential orientation of (002) planes. *Environ. Sci. Technol.* **2007**, *41*, 4422–4427,
517 doi:10.1021/es062546c.
- 518 14. Bamwenda, G.R.; Arakawa, H. Visible light induced photocatalytic activity of tungsten trioxide
519 powders. *Appl. Catal. A Gen.* **2001**, *210*, 181–191, doi:10.1016/S0926-860X(00)00796-1.
- 520 15. Gondal, M.A.; Dastageer, M.A.; Khalil, A. Synthesis of nano-WO₃ and its catalytic activity for enhanced
521 antimicrobial process for water purification using laser induced photo-catalysis. *Catal. Commun.* **2009**,
522 *11*, 214–219, doi:10.1016/j.catcom.2009.10.011.
- 523 16. Waldner, G.; Brüger, A.; Gaikwad, N.S.; Neumann-Spallart, M. WO₃ thin films for photoelectrochemical
524 purification of water. *Chemosphere* **2007**, *67*, 779–784, doi:10.1016/j.chemosphere.2006.10.024.
- 525 17. Su, X.; Li, Y.; Jian, J.; Wang, J. In situ etching WO₃ nanoplates: Hydrothermal synthesis,

- 526 photoluminescence and gas sensor properties. *Mater. Res. Bull.* **2010**, *45*, 1960–1963,
527 doi:10.1016/j.materresbull.2010.08.011.
- 528 18. An, S.; Park, S.; Ko, H.; Lee, C. Fabrication of WO₃ nanotube sensors and their gas sensing properties.
529 *Ceram. Int.* **2014**, *40*, 1423–1429, doi:10.1016/J.CERAMINT.2013.07.025.
- 530 19. Li, Z.; Li, J.; Song, L.; Gong, H.; Niu, Q. Ionic liquid-assisted synthesis of WO₃ particles with enhanced
531 gas sensing properties. *J. Mater. Chem. A* **2013**, *1*, 15377, doi:10.1039/c3ta13500d.
- 532 20. Miseki, Y.; Kusama, H.; Sugihara, H.; Sayama, K. Cs-modified WO₃ photocatalyst showing efficient solar
533 energy conversion for O₂ production and Fe (III) ion reduction under visible light. *J. Phys. Chem. Lett.*
534 **2010**, *1*, 1196–1200, doi:10.1021/jz100233w.
- 535 21. Miseki, Y.; Kusama, H.; Sayama, K. Photocatalytic energy storage over surface-modified WO₃ using V
536 ⁵⁺/V⁴⁺ redox mediator. *Chem. Lett.* **2012**, *41*, 1489–1491, doi:10.1246/cl.2012.1489.
- 537 22. Choi, Y.G.; Sakai, G.; Shimano, K.; Teraoka, Y.; Miura, N.; Yamazoe, N. Preparation of size and habit-
538 controlled nano crystallites of tungsten oxide. *Sensors Actuators, B Chem.* **2003**, *93*, 486–494,
539 doi:10.1016/S0925-4005(03)00195-3.
- 540 23. Santato, C.; Odziemkowski, M.; Ulmann, M.; Augustynski, J. Crystallographically oriented mesoporous
541 WO₃ films: Synthesis, characterization, and applications. *J. Am. Chem. Soc.* **2001**, *123*, 10639–10649,
542 doi:10.1021/ja011315x.
- 543 24. Badilescu, S.; Ashrit, P. V. Study of sol-gel prepared nanostructured WO₃ thin films and composites for
544 electrochromic applications. *Solid State Ionics* **2003**, *158*, 187–197, doi:10.1016/S0167-2738(02)00764-6.
- 545 25. Yu, Z.; Jia, X.; Du, J.; Zhang, J. Electrochromic WO₃ films prepared by a new electrodeposition method.
546 *Sol. Energy Mater. Sol. Cells* **2000**, *64*.
- 547 26. Brescacin, E.; Basato, M.; Tondello, E. Amorphous WO₃ films via chemical vapor deposition from
548 metallorganic precursors containing phosphorus dopant. *Chem. Mater.* **1999**, *11*, 314–323,
549 doi:10.1021/cm980741n.
- 550 27. Palgrave, R.G.; Parkin, I.P. Aerosol assisted chemical vapour deposition of photochromic tungsten oxide
551 and doped tungsten oxide thin films. *J. Mater. Chem.* **2004**, *14*, 2864–2867, doi:10.1039/b406337f.
- 552 28. Tsuchiya, H.; Macak, J.M.; Sieber, I.; Taveira, L.; Ghicov, A.; Sirotna, K.; Schmuki, P. Self-organized
553 porous WO₃ formed in NaF electrolytes. *Electrochem. commun.* **2005**, *7*, 295–298,
554 doi:10.1016/j.elecom.2005.01.003.
- 555 29. Ng, C.; Ye, C.; Ng, Y.H.; Amal, R. Flower-shaped tungsten oxide with inorganic fullerene-like structure:
556 Synthesis and characterization. *Cryst. Growth Des.* **2010**, *10*, 3794–3801, doi:10.1021/cg100625m.
- 557 30. Srichaiyaperk, T.; Aiempanakit, K.; Horprathum, M.; Eiamchai, P.; Chananonawathorn, C.;
558 Limwichean, S.; Chindaudom, P. Effects of annealing treatment on WO₃ thin films prepared by DC
559 reactive magnetron sputtering. *Adv. Mater. Res.* **2014**, *979*, 248–250,
560 doi:10.4028/www.scientific.net/AMR.979.248.
- 561 31. Liang, Y.C.; Chang, C.W. Preparation of orthorhombic WO₃ thin films and their crystal quality-
562 dependent dye photodegradation ability. *Coatings* **2019**, *9*, doi:10.3390/COATINGS9020090.
- 563 32. Marszalek, K. Magnetron-sputtered WO₃ films for electrochromic devices. *Thin Solid Films* **1989**, *175*,
564 227–233, doi:10.1016/0040-6090(89)90832-8.
- 565 33. M., B.B.; K., V.M. Structural, morphological and optical properties of electron beam evaporated WO₃
566 thin films. *J. Taibah Univ. Sci.* **2017**, *11*, 1232–1237, doi:10.1016/j.jtusc.2016.12.003.
- 567 34. Sivakumar, R.; Gopalakrishnan, R.; Jayachandran, M.; Sanjeeviraja, C. Preparation and characterization
568 of electron beam evaporated WO₃ thin films. *Opt. Mater. (Amst)*. **2007**, *29*, 679–687,

- 569 doi:10.1016/j.optmat.2005.11.017.
- 570 35. Malm, J.; Sajavaara, T.; Karppinen, M. Atomic layer deposition of WO₃ thin films using W(CO)₆ and O₃
- 571 precursors. *Chem. Vap. Depos.* **2012**, *18*, 245–248, doi:10.1002/cvde.201206986.
- 572 36. Li, B.; Gao, X.; Zhang, H.C.; Yuan, C. Energy modeling of electrochemical anodization process of
- 573 titanium dioxide nanotubes. *ACS Sustain. Chem. Eng.* **2014**, *2*, 404–410, doi:10.1021/sc400304h.
- 574 37. Berger, S.; Tsuchiya, H.; Ghicov, A.; Schmuki, P. High photocurrent conversion efficiency in self-
- 575 organized porous WO₃. *Appl. Phys. Lett.* **2006**, *88*, 88–91, doi:10.1063/1.2206696.
- 576 38. Caramori, S.; Cristino, V.; Meda, L.; Tacca, A.; Argazzi, R.; Bignozzi, C.A. Efficient anodically grown
- 577 WO₃ for photoelectrochemical water splitting. *Energy Procedia* **2012**, *22*, 127–136,
- 578 doi:10.1016/j.egypro.2012.05.214.
- 579 39. Nah, Y.C.; Ghicov, A.; Kim, D.; Schmuki, P. Enhanced electrochromic properties of self-organized
- 580 nanoporous WO₃. *Electrochem. Commun.* **2008**, *10*, 1777–1780, doi:10.1016/j.elecom.2008.09.017.
- 581 40. Ou, J.Z.; Rani, R.A.; Balendhran, S.; Zoofakar, A.S.; Field, M.R.; Zhuiykov, S.; O'Mullane, A.P.; Kalantar-
- 582 Zadeh, K. Anodic formation of a thick three-dimensional nanoporous WO₃ film and its photocatalytic
- 583 property. *Electrochem. Commun.* **2013**, *27*, 128–132, doi:10.1016/j.elecom.2012.11.009.
- 584 41. Ou, J.Z.; Balendhran, S.; Field, M.R.; McCulloch, D.G.; Zoofakar, A.S.; Rani, R.A.; Zhuiykov, S.;
- 585 O'Mullane, A.P.; Kalantar-Zadeh, K. The anodized crystalline WO₃ nanoporous network with enhanced
- 586 electrochromic properties. *Nanoscale* **2012**, *4*, 5980–5988, doi:10.1039/c2nr31203d.
- 587 42. Mukherjee, N.; Paulose, M.; Varghese, O.K.; Mor, G.K.; Grimes, C.A. Fabrication of nanoporous tungsten
- 588 oxide by galvanostatic anodization. *J. Mater. Res.* **2003**, *18*, 2296–2299, doi:10.1557/JMR.2003.0321.
- 589 43. Yu, L.; Kalantar-zadeh, K.; Zhu, J.; Latham, K.; Wlodarski, W.; Hu, Z.; Sadek, A.Z.; Breedon, M.; Zheng,
- 590 H.; Bhargava, S.K.; et al. High-Temperature Anodized WO₃ Nanoplatelet Films for Photosensitive
- 591 Devices. *Langmuir* **2009**, *25*, 9545–9551, doi:10.1021/la901944x.
- 592 44. Ng, C.; Ng, Y.H.; Iwase, A.; Amal, R. Influence of annealing temperature of WO₃ in photoelectrochemical
- 593 conversion and energy storage for water splitting. *ACS Appl. Mater. Interfaces* **2013**, *5*, 5269–5275,
- 594 doi:10.1021/am401112q.
- 595 45. Nischk, M.; Mazierski, P.; Gazda, M.; Zaleska, A. Ordered TiO₂ nanotubes: The effect of preparation
- 596 parameters on the photocatalytic activity in air purification process. *Appl. Catal. B Environ.* **2014**, *144*,
- 597 674–685, doi:10.1016/j.apcatb.2013.07.041.
- 598 46. Macak, J.M.; Schmuki, P. Anodic growth of self-organized anodic TiO₂ nanotubes in viscous electrolytes.
- 599 *Electrochim. Acta* **2006**, *52*, 1258–1264, doi:10.1016/j.electacta.2006.07.021.
- 600 47. Pourbaix, M. Atlas of electrochemical equilibria in aqueous solutions 2 English ed. *Natl. Assoc. Corros.*
- 601 *Eng. Houston, TX*, **1974**, doi:10.1016/0022-0728(67)80059-7.
- 602 48. Roy, P.; Berger, S.; Schmuki, P. TiO₂ nanotubes: Synthesis and applications. *Angew. Chemie - Int. Ed.* **2011**,
- 603 50, 2904–2939, doi:10.1002/anie.201001374.
- 604 49. Macak, J.M.; Tsuchiya, H.; Ghicov, A.; Yasuda, K.; Hahn, R.; Bauer, S.; Schmuki, P. TiO₂ nanotubes: Self-
- 605 organized electrochemical formation, properties and applications. *Curr. Opin. Solid State Mater. Sci.* **2007**,
- 606 11, 3–18, doi:10.1016/J.COSSMS.2007.08.004.
- 607 50. Chai, Y.; Tam, C.W.; Beh, K.P.; Yam, F.K.; Hassan, Z. Porous WO₃ formed by anodization in oxalic acid.
- 608 *J. Porous Mater.* **2013**, *20*, 997–1002, doi:10.1007/s10934-013-9675-5.
- 609 51. Anik, M.; Osseo-Asare, K. Effect of pH on the Anodic Behavior of Tungsten. *J. Electrochem. Soc.* **2002**, *149*,
- 610 B224, doi:10.1149/1.1471544.
- 611 52. Macdonald, D.D.; Sikora, E.; Sikora, J. The kinetics of growth of the passive film on tungsten in acidic

- 612 phosphate solutions. *Electrochim. Acta* **1998**, *43*, 2851–2861, doi:10.1016/S0013-4686(98)00026-7.
- 613 53. Lai, C.W. Photocatalysis and photoelectrochemical properties of tungsten trioxide nanostructured films.
- 614 *Sci. World J.* **2014**, *2014*, doi:10.1155/2014/843587.
- 615 54. El-Basiouny, M.S.; Hassan, S.A.; Hefny, M.M. On the electrochemical behaviour of tungsten: the
- 616 formation and dissolution of tungsten oxide in sulphuric acid solutions. *Corros. Sci.* **1980**, *20*, 909–917,
- 617 doi:10.1016/0010-938X(80)90123-7.
- 618 55. Baserga, A.; Russo, V.; Di Fonzo, F.; Bailini, A.; Cattaneo, D.; Casari, C.S.; Li Bassi, A.; Bottani, C.E.
- 619 Nanostructured tungsten oxide with controlled properties: Synthesis and Raman characterization. *Thin*
- 620 *Solid Films* **2007**, *515*, 6465–6469, doi:10.1016/j.tsf.2006.11.067.
- 621 56. Daniel, M.F.; Desbat, B.; Lassegues, J.C.; Gerand, B.; Figlarz, M. Infrared and Raman study of WO₃
- 622 tungsten trioxides and WO₃·xH₂O tungsten trioxide hydrates. *J. Solid State Chem.* **1987**, *67*, 235–247,
- 623 doi:10.1016/0022-4596(87)90359-8.
- 624 57. Butler, M.A. Photoelectrolysis and physical properties of the semiconducting electrode WO₂. *J. Appl.*
- 625 *Phys.* **1977**, *48*, 1914–1920, doi:10.1063/1.323948.
- 626 58. Teoh, L.G.; Shieh, J.; Lai, W.H.; Hung, I.M.; Hon, M.H. Structure and optical properties of mesoporous
- 627 tungsten oxide. *J. Alloys Compd.* **2005**, *396*, 251–254, doi:10.1016/j.jallcom.2004.12.005.
- 628 59. Fujihara, S.; Ogawa, Y.; Kasai, A. Tunable visible photoluminescence from ZnO thin films through Mg-
- 629 doping and annealing. *Chem. Mater.* **2004**, *16*, 2965–2968, doi:10.1021/cm049599i.
- 630 60. Sreekantan, S.; Saharudin, K.A.; Lockman, Z.; Tzu, T.W. Fast-rate formation of TiO₂ nanotube arrays in
- 631 an organic bath and their applications in photocatalysis. *Nanotechnology* **2010**, *21*, doi:10.1088/0957-
- 632 4484/21/36/365603.
- 633 61. Liang, Y.; Lin, S.; Liu, L.; Hu, J.; Cui, W. Oil-in-water self-assembled Ag@AgCl QDs sensitized Bi₂WO₆:
- 634 Enhanced photocatalytic degradation under visible light irradiation. *Appl. Catal. B Environ.* **2015**, *164*,
- 635 192–203, doi:10.1016/j.apcatb.2014.08.048.
- 636 62. Zhu, Z.; Yan, Y.; Li, J. Preparation of flower-like BiOBr-WO₃-Bi₂WO₆ ternary hybrid with enhanced
- 637 visible-light photocatalytic activity. *J. Alloys Compd.* **2015**, *651*, 184–192, doi:10.1016/j.jallcom.2015.08.137.
- 638 63. Kalantar-zadeh, K.; Sadek, A.Z.; Zheng, H.; Bansal, V.; Bhargava, S.K.; Wlodarski, W.; Zhu, J.; Yu, L.;
- 639 Hu, Z. Nanostructured WO₃ films using high temperature anodization. *Sensors Actuators, B Chem.* **2009**,
- 640 *142*, 230–235, doi:10.1016/j.snb.2009.08.014.
- 641 64. Zhuang, H.; Lin, C.; Lai, Y.; Sun, L.; Li, J. Some Critical Structure Factors of Titanium Oxide Nanotube
- 642 Array in Its Photocatalytic Activity Some Critical Structure Factors of Titanium Oxide Nanotube Array
- 643 in Its Photocatalytic Activity. **2007**, 4735–4740, doi:10.1021/es0702723.
- 644



© 2020 by the authors. Submitted for possible open access publication under the terms and conditions of the Creative Commons Attribution (CC BY) license (<http://creativecommons.org/licenses/by/4.0/>).